

### 1.2.1 Compare (qualitatively) the relative strengths of ionic, covalent, and metallic bonds.

1. Compare the relative strengths of the three different types of bonds.

ionic > covalent > metallic

2. Which statement compares the amount of energy needed to break the bonds in  $\text{CaCl}_2$  (E1) and  $\text{C}_{12}\text{H}_{22}\text{O}_{11}$  (E2)?

- a.  $\text{E1} > \text{E2}$ , as  $\text{CaCl}_2$  is a covalent compound.
- b.  $\text{E1} < \text{E2}$ , as  $\text{CaCl}_2$  is a covalent compound.
- ☒ c.  $\text{E1} > \text{E2}$ , as  $\text{CaCl}_2$  is an ionic compound.
- d.  $\text{E1} < \text{E2}$ , as  $\text{CaCl}_2$  is an ionic compound.

### 1.2.2 Infer the type of bond and chemical formula formed between atoms.

1. Which pair of elements would *most likely* bond to form a covalently bonded compound?

- a. sodium and fluorine
- b. barium and chlorine
- ☒ c. phosphorus and oxygen
- d. magnesium and sulfur

2. For each pair of atoms, predict whether the bond formed between the atoms is either ionic or covalent, and write the formula for the predicted compound.

- 1. Na and O: ionic,  $\text{Na}_2\text{O}$
- 2. S and F: covalent,  $\text{SF}_2$
- 3. Ag and N: ionic,  $\text{Ag}_3\text{N}$
- 4. Te and H: covalent,  $\text{TeH}_2$

ionic = M + NM (or polyatomic ion)  
covalent = all NM  
metallic = all M

3. Which statement describes the compound formed between sodium and oxygen?

- a. It is  $\text{NaO}_2$ , which is ionic.
- b. It is  $\text{NaO}_2$ , which is covalent.
- ☒ c. It is  $\text{Na}_2\text{O}$ , which is ionic.
- d. It is  $\text{Na}_2\text{O}$ , which is covalent.

### 1.2.3 Compare inter- and intra- particle forces.

1. Rank the following substances in the order in which they would evaporate, justifying the order of placement for each (using those inter- and intra- particle forces).

- 1. Water ( $\text{H}_2\text{O}$ )
- 2. Methane ( $\text{CH}_4$ )
- 3. Sodium chloride ( $\text{NaCl}$ )
- 4. Phosphorus trifluoride ( $\text{PF}_3$ )

1.  $\text{CH}_4$  - LDFs (weakest)  
2.  $\text{PF}_3$  - LDFs, dipole-dipole  
3.  $\text{H}_2\text{O}$  - LDFs, dipole-dipole, H-bonding  
4.  $\text{NaCl}$  - ionic bond (strongest forces)

2. At STP, fluorine is a gas and iodine is a solid. Why?

- a. Fluorine has lower average kinetic energy than iodine.
- b. Fluorine has higher average kinetic energy than iodine.
- ☒ c. Fluorine has weaker intermolecular forces of attraction than iodine.
- d. Fluorine has stronger intermolecular forces of attraction than iodine.

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**TD Snyder**



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## **Table of Contents Math Skills Transparency Worksheet Answers Chapter 6**

1. Understanding the eBook Math Skills Transparency Worksheet Answers Chapter 6
  - The Rise of Digital Reading Math Skills Transparency Worksheet Answers Chapter 6
  - Advantages of eBooks Over Traditional Books
2. Identifying Math Skills Transparency Worksheet Answers Chapter 6
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Math Skills Transparency Worksheet Answers Chapter 6
  - User-Friendly Interface
4. Exploring eBook Recommendations from Math Skills Transparency Worksheet Answers Chapter 6
  - Personalized Recommendations
  - Math Skills Transparency Worksheet Answers Chapter 6 User Reviews and Ratings
  - Math Skills Transparency Worksheet Answers Chapter 6 and Bestseller Lists

5. Accessing Math Skills Transparency Worksheet Answers Chapter 6 Free and Paid eBooks
  - Math Skills Transparency Worksheet Answers Chapter 6 Public Domain eBooks
  - Math Skills Transparency Worksheet Answers Chapter 6 eBook Subscription Services
  - Math Skills Transparency Worksheet Answers Chapter 6 Budget-Friendly Options
6. Navigating Math Skills Transparency Worksheet Answers Chapter 6 eBook Formats
  - ePub, PDF, MOBI, and More
  - Math Skills Transparency Worksheet Answers Chapter 6 Compatibility with Devices
  - Math Skills Transparency Worksheet Answers Chapter 6 Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Math Skills Transparency Worksheet Answers Chapter 6
  - Highlighting and Note-Taking Math Skills Transparency Worksheet Answers Chapter 6
  - Interactive Elements Math Skills Transparency Worksheet Answers Chapter 6
8. Staying Engaged with Math Skills Transparency Worksheet Answers Chapter 6
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Math Skills Transparency Worksheet Answers Chapter 6
9. Balancing eBooks and Physical Books Math Skills Transparency Worksheet Answers Chapter 6
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Math Skills Transparency Worksheet Answers Chapter 6
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Math Skills Transparency Worksheet Answers Chapter 6
  - Setting Reading Goals Math Skills Transparency Worksheet Answers Chapter 6
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Math Skills Transparency Worksheet Answers Chapter 6
  - Fact-Checking eBook Content of Math Skills Transparency Worksheet Answers Chapter 6
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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