

Name \_\_\_\_\_

Chemistry Period \_\_\_\_\_

### Worksheet Specific Heat

Specific Heat of water =  $4.18 \text{ J/g}^\circ\text{C}$

1. A sample of mercury is heated from  $25.5^\circ\text{C}$  to  $52.5^\circ\text{C}$ . In the process,  $3050\text{J}$  of heat are absorbed. What mass of mercury was in the sample? The specific heat of mercury is  $0.140 \text{ J/g}^\circ\text{C}$ . **ANS: 807 g**
2. A block of aluminum with a mass of  $140\text{g}$  is cooled from  $98.4^\circ\text{C}$  to  $62.2^\circ\text{C}$  with a release of  $1137\text{J}$  of heat. From these data, calculate the specific heat of aluminum. **ANS: 0.224**
3. A total of  $226 \text{ J}$  of heat are absorbed as  $58.3 \text{ g}$  of lead is heated from  $12.0^\circ\text{C}$  to  $42.0^\circ\text{C}$ . From this data, what is the specific heat of lead? **ANS: 0.129**
4. What is the change in heat energy when  $132\text{g}$  of copper at  $32.2^\circ\text{C}$  is raised to  $45.0^\circ\text{C}$ ? The specific heat of copper is  $0.385 \text{ J/g}^\circ\text{C}$ . **ANS: 650 J**
5. What is the change in heat energy when  $40.0\text{g}$  of water at  $25.0^\circ\text{C}$  is raised to  $60.0^\circ\text{C}$ ? **ANS: 5852 J**
6. What is the change in heat energy when  $75.0$  of water at  $9.8^\circ\text{C}$  is raised to  $22.4^\circ\text{C}$ ? **3950 J**
7. What is the change in heat energy when  $125\text{g}$  of water at  $10.0^\circ\text{C}$  is raised to  $90.0^\circ\text{C}$ ? **ANS: 41.8 kJ**
8. What is the change in heat energy when  $64.82 \text{ g}$  of aluminum metal at  $100.0^\circ\text{C}$  is cooled to  $82.0^\circ\text{C}$ ? The specific heat of aluminum is  $0.897 \text{ J/g}^\circ\text{C}$ . **ANS: 1047**
9. What is the change in temperature if  $24\text{g}$  of water at  $50^\circ\text{C}$  absorbs  $985 \text{ J}$  of heat energy? **ANS: 9.82**
10. What is the mass of a sample of iron if that sample lost  $2300\text{J}$  of heat energy when it cooled from  $80^\circ\text{C}$  to  $30^\circ\text{C}$ ? The specific heat of iron is  $0.449 \text{ J/g}^\circ\text{C}$ . **ANS: 102 g**
11. What is the specific heat of metal if its mass is  $26.86\text{g}$  and it requires  $418.6 \text{ J}$  of heat energy to raise its temperature from  $27.4^\circ\text{C}$  to  $67.3^\circ\text{C}$ ? **ANS: 0.391**
12. How much heat energy is needed to raise the temperature of  $59.7\text{g}$  of cadmium from  $25^\circ\text{C}$  to  $100^\circ\text{C}$ ? The specific heat of cadmium is  $0.231 \text{ J/g}^\circ\text{C}$ . **ANS: 1034 J**

# Math Skills Specific Heat Answers

**Sebastian Brünink**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

## **Math Skills Specific Heat Answers:**

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Kids, Carson-Dellosa Publishing, 2016-03-07 Comprehensive Curriculum of Basic Skills for grade 4 covers basic concepts such as reading comprehension parts of speech punctuation spelling paragraphs research multiplication division estimating averaging fractions decimals graphing geometry and probability Complete with practice in writing reading and math this series helps develop the skills your child needs for grade level success With over 10 million copies in print the Comprehensive Curriculum of Basic Skills series provides an entire curriculum filled with fun educational activities and instruction that improve academic performance Available for grades prekindergarten to 6 Comprehensive Curriculum of Basic Skills features vivid full color illustrations and grade appropriate activities for phonics reading language arts writing and math This series edition has been updated with relevant high interest reading passages and artwork to engage your child in the learning process An excellent resource for supporting classroom learning or enhancing your home school curriculum it features review lessons to measure your child's progress teaching suggestions to extend learning and answer keys to monitor accuracy Comprehensive Curriculum of Basic Skills is the all in one resource for strengthening essential skills

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