

The Significance of Electrons in the Transportation of Energy in Living Systems

In biological systems, Oxidation and Reduction Reactions are critical in the production and transfer of energy. Usually, an Oxidation Reaction removes electrons from one compound and transfer same to another compound. This addition of electrons to another compound is called Reduction Reaction. These two reactions occur simultaneously, and is referred to as Redox Reaction (Libretexts, 2012).

The transfer of electrons between compounds or molecules is important. This is because most of the energy stored in atoms and deploy in fueling cell functions is normally "in the form of high-energy electrons". When this high energy is transferred in the form of electrons, it enables the cell to transport and utilize the energy in an incremental manner rather than in a single, destructive fashion (Libretexts, 2012).

Major carriers of electron are derived from Vitamin B₃ group, and they are the derivatives of nucleotides, and they are easily reducible and oxidizing compounds (Libretexts, 2012).

Typical example includes Nicotinamide Adenine Dinucleotide, which is abbreviated as NAD⁺. The NAD⁺ has two rings, and one of them is the nicotinamide ring, which is the active part of the molecule that accepts and donates two electrons. Nicotinamide Adenine Dinucleotide serves as an electron acceptor during the process of glycolysis, and the cycle of citric acid of cellular respiration and donates same to oxidative phosphorylation (Becker, n.d.).

In conclusion, electrons are vital in the transportation of energy between molecules.

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