



# **Maintenance & Service Guide**

## **Sprout Business PC**

# Hp M7334 Desktops Owners Manual

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web start with 8.3 mole  $\text{H}_2\text{S}$  and multiply it with 2 mole  $\text{H}_2\text{O}$  over 2 mole  $\text{H}_2\text{S}$  from the equation  $8.3\text{ mole H}_2\text{S} \times \frac{2\text{ mole H}_2\text{O}}{2\text{ mole H}_2\text{S}} \times 18\text{g H}_2\text{O} = 149.4\text{g H}_2\text{O}$  is your theoretical yield so percent yield  $\frac{137.1\text{g H}_2\text{O}}{149.4\text{g H}_2\text{O}} \times 100 = 91.77\%$  problems 6 & 8 are similar to problem 5

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web a write a balanced equation for the reaction  $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$  4 Al 3  $\text{O}_2$  2  $\text{Al}_2\text{O}_3$  b determine the theoretical yield of  $\text{Al}_2\text{O}_3$  160.0 g Al  $\times \frac{1 \text{ mol Al}}{2 \text{ mol Al}_2\text{O}_3} \times \frac{101.964 \text{ g Al}_2\text{O}_3}{3 \text{ g Al}_2\text{O}_3} = 26.982 \text{ g Al}_2\text{O}_3$  c determine the percent yield  $\frac{\text{actual}}{\text{theoretical}} \times 100 = \frac{26.0 \text{ g}}{26.982 \text{ g}} \times 100 = 96.01\%$

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web 7 Na 2 C 2 O 4 is the limiting reactant percent yield 86.6% only four molecules can be made 9 this amount cannot be weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12 C is the limiting reagent 4.33 g of H 2 are left over 13

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