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Chapter 17: Thermochemistry

Section 1: The Flow of Energy - Heat and Work

- What is thermochemistry? The study of the energy changes that occur in changes of state or during chemical reactions.
- 2. Heat always flows from a warmer object to a cooler object.
- Given that 1 cal = 4.184 J:
 - a. How many calories are in 325.9 J? 77.89 cal
 - How many J are in 34.7 cal? 145.18 J
- What is the Law of Conservation of Energy (aka the First Law of Thermodynamics)? States that in chemical and physical processes, energy is neither created nor destroyed – it is transferred or conserved.
- 5. In an exothermic reaction:
 - a. Heat is released / absorbed
 - b. q is positive / negative
 - c. ΔH is positive / negative
 - d. Heat is a reactant / product
- 6. In an endothermic reaction:
 - a. Heat is released / absorbed
 - b. q is positive / negative
 - c. ΔH is positive / negative
 - d. Heat is a reactant / product
- Diagram A represents an exothermic reaction, whereas diagram B represents an endothermic reaction:
- What is the difference between heat capacity and specific heat capacity? Specific heat is the amount of energy to increase the temperature of <u>1g</u> of substance by 1°C
- The temperature of a 34.1 g sample of a substance increases from 29 C to 51 C when the sample absorbs 912 J of heat. What is the specific heat of the substance? 1.126 J/g*C
- 10. When 867 J of heat are added to 9.3 g of a substance at STP, the temperature increases from 0 C to 14 C. What is the specific heat of the substance? 6.66 J/g C

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