

Radioactivity and Half-Lives Review Worksheet

1. What is the difference between radioactivity and nuclear radiation?

Radioactivity is the process of nuclear decay (an unstable isotope emitting energy/a particle to become more stable). Nuclear radiation is that energy or particle being emitted.

2. Which type of radiation can be blocked by---

- a. A piece of paper? **alpha**
- b. A piece of wood? **beta**
- c. A large block of concrete? **gamma**

Extension opportunity: Why? *What do we know about the size of each type of particle that affects their ability to pass through materials?* **Alpha particles are the largest, so they move the slowest and cannot penetrate objects as easily. Gamma particles are the smallest and have no charge, so they penetrate more easily.**

3. What is a half-life?

The time it takes for half of the material you started with to decay.

4. If we start with 400 atoms of a radioactive substance, how many would remain after one half-life? **200** After 2 half-lives? **100** After 3 half-lives? **50** After 4 half-lives? **25**

5. If we start with 16 grams of a substance, how much will remain after 3 half-lives?
 $16 \rightarrow 8 \rightarrow 4 \rightarrow 2$ **2 grams** after 3 half-lives.

6. A sample of carbon-14 has been decaying for 22,920 years and is now 35.0 grams. What was the size of the original sample? (The half-life of carbon-14 is 5,730 years)
 $22920/5730 = 4$ half lives *You will need to work backwards*
 $35 \rightarrow 70 \rightarrow 140 \rightarrow 280 \rightarrow$ **560. grams**

7. How old is a skeleton sample if the current amount of carbon in the bones is 3.125%? (Assume you started with 100%)
 $100\% \rightarrow 50 \rightarrow 25 \rightarrow 12.5 \rightarrow 6.25 \rightarrow 3.125$ 5 half lives $\times 5,730 =$ **28,650 years**

8. If you started with 120 grams of a radioactive substance, and now you have 15 grams, how many half-lives have passed?
 $120 \rightarrow 60 \rightarrow 30 \rightarrow 15$ **3 half-lives**

9. The half-life of isotope X is 2.0 years. How many years would it take for a 4.0 mg sample of isotope X to decay and have only 0.50 mg of it remain?
 $4.0 \rightarrow 2.0 \rightarrow 1.0 \rightarrow 0.50$ 3 half lives $\times 2$ years = **6.0 years**

Half Life Calculations Worksheet Answer Key

Ying-Ying Zheng

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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Table of Contents Half Life Calculations Worksheet Answer Key

1. Understanding the eBook Half Life Calculations Worksheet Answer Key
 - The Rise of Digital Reading Half Life Calculations Worksheet Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Half Life Calculations Worksheet Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Half Life Calculations Worksheet Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Half Life Calculations Worksheet Answer Key

- Personalized Recommendations
 - Half Life Calculations Worksheet Answer Key User Reviews and Ratings
 - Half Life Calculations Worksheet Answer Key and Bestseller Lists
5. Accessing Half Life Calculations Worksheet Answer Key Free and Paid eBooks
 - Half Life Calculations Worksheet Answer Key Public Domain eBooks
 - Half Life Calculations Worksheet Answer Key eBook Subscription Services
 - Half Life Calculations Worksheet Answer Key Budget-Friendly Options
 6. Navigating Half Life Calculations Worksheet Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Half Life Calculations Worksheet Answer Key Compatibility with Devices
 - Half Life Calculations Worksheet Answer Key Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Half Life Calculations Worksheet Answer Key
 - Highlighting and Note-Taking Half Life Calculations Worksheet Answer Key
 - Interactive Elements Half Life Calculations Worksheet Answer Key
 8. Staying Engaged with Half Life Calculations Worksheet Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Half Life Calculations Worksheet Answer Key
 9. Balancing eBooks and Physical Books Half Life Calculations Worksheet Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Half Life Calculations Worksheet Answer Key
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Half Life Calculations Worksheet Answer Key
 - Setting Reading Goals Half Life Calculations Worksheet Answer Key
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Half Life Calculations Worksheet Answer Key

- Fact-Checking eBook Content of Half Life Calculations Worksheet Answer Key
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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