

Empirical and Molecular Formulas Practice

Read the question carefully for the data provided. Answers must be shown for credit.

Find empirical formula:

1. A compound is found to contain 40.0 g of carbon and 6.67 g of hydrogen. Find the empirical formula.

$$\text{C}_{40.0} \text{H}_{6.67} \times \frac{1 \text{ mole}}{12.01 \text{ g C}} = 3.33 \text{ moles C}$$

$$\text{H}_{6.67} \times \frac{1 \text{ mole}}{1.008 \text{ g H}} = 6.61 \text{ moles H}$$

C_{3.33}H_{6.61}

2. A compound is found to contain 50.0 g of carbon, 10.0 g of hydrogen, and 33.3 g of oxygen. Find the empirical formula.

$$\text{C}_{50.0} \text{H}_{10.0} \text{O}_{33.3} \times \frac{1 \text{ mole}}{12.01 \text{ g C}} = 4.16 \text{ moles C}$$

$$\text{H}_{10.0} \times \frac{1 \text{ mole}}{1.008 \text{ g H}} = 9.92 \text{ moles H}$$



$$\text{O}_{33.3} \times \frac{1 \text{ mole}}{16.00 \text{ g O}} = 2.08 \text{ moles O} \quad \text{C}_{4.16} \text{H}_{9.92} \text{O}_{2.08} \rightarrow \text{C}_4\text{H}_{10}\text{O}_2$$

3. A 10.0 g sample of a compound, a gas, contains 2.50 g of carbon, 0.50 g of hydrogen, and 7.00 g of oxygen. Find the empirical formula.

$$\text{C}_{2.50} \text{H}_{0.50} \text{O}_{7.00} \times \frac{1 \text{ mole}}{12.01 \text{ g C}} = 0.208 \text{ moles C}$$



$$\text{H}_{0.50} \times \frac{1 \text{ mole}}{1.008 \text{ g H}} = 0.496 \text{ moles H} \quad \text{C}_{0.208} \text{H}_{0.496} \text{O}_{3.37} \rightarrow \text{C}_2\text{H}_5\text{O}_{16}$$

$$\text{O}_{7.00} \times \frac{1 \text{ mole}}{16.00 \text{ g O}} = 0.438 \text{ moles O} \quad \text{C}_{0.208} \text{H}_{0.496} \text{O}_{3.37} \rightarrow \text{C}_2\text{H}_5\text{O}_{16}$$

4. A 20.00 g sample of an unknown compound contains 12.00 g carbon, 2.00 g hydrogen, and 6.00 g oxygen. Find the empirical formula.

$$\text{C}_{12.00} \text{H}_{2.00} \text{O}_{6.00} \times \frac{1 \text{ mole}}{12.01 \text{ g C}} = 1.00 \text{ moles C}$$



$$\text{H}_{2.00} \times \frac{1 \text{ mole}}{1.008 \text{ g H}} = 1.98 \text{ moles H}$$

$$\text{O}_{6.00} \times \frac{1 \text{ mole}}{16.00 \text{ g O}} = 0.375 \text{ moles O} \quad \text{C}_{1.00} \text{H}_{1.98} \text{O}_{0.375} \rightarrow \text{C}_8\text{H}_{16}\text{O}_3$$

Empirical Formula Pogil

JG Myers



Empirical Formula Pogil:

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