

Empirical and Molecular Formula Practice

3 pages + 3-page Answer KEY + Periodic Table

EMPIRICAL AND MOLECULAR PROBLEMS

Chemistry

1 The percent composition of a certain gas is nitrogen (28.1%) and hydrogen (7.14%). Calculate the empirical formula.
Show & Convert it to %

Step 1: Convert % to g/mol

Step 3: Divide by the smallest mole value

Step 4: Multiply by smallest number that will make a whole

2 The compound in problem 1 has a molar mass of 40 g/mol.
Show 1. Calculate the molar mass of the empirical formula

Step 2: Divide molecular mass (molar) by empirical mass

Step 5: Multiply empirical formula by the answer in Step 4

3 A compound is composed of 80.1% nitrogen and 19.9% hydrogen. What is the empirical formula of this compound?

4 The compound in problem 3 has a molar mass of 100 g/mol.

5 Hydrogen's molar mass is 1 g/mol. Calculate the empirical formula

6 Hydrogen's molar mass is 1 g/mol

7 A compound composed of hydrogen (2.16%) and oxygen (77.8%), has a molar mass of 34 g/mol. What is the empirical formula of this compound?

8 What is the molecular formula of the compound in problem 7?

9 A compound with a known molar mass of 34 g/mol contains carbon (24.04%) and hydrogen (75.96%). What is the empirical formula?

10 Calculate the molecular formula of the compound in problem 9

ANSWER KEY
INCLUDED

Empirical And Molecular Calculation Practice Packet

Answer Key

Lei Shi



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number to save room the configurations are in noble gas shorthand this means part of the electron configuration has been replaced with the element symbol of the noble gas symbol

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web elements in the second row of the periodic table place their electrons in the 2n shell as well as the 1n shell for instance lithium li has three electrons two fill the 1 s orbital and the third is placed in the 2 s orbital giving an electron configuration of 1 s 2 2 s 1

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web the electron removed during the ionization of beryllium he 2s² 2 is an s electron whereas the electron removed during the ionization of boron he 2s² 2p¹ is a p electron this results in a lower first ionization energy for boron even though its nuclear charge is greater by one proton

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web predictions from reliable sources have been used for these elements grayed out electron numbers indicate subshells filled to their maximum bracketed noble gas symbols on the left represent inner configurations that are the same in each period written out these are he 2 helium 1s² ne 10 neon 1s² 2s² 2p⁶

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