

Cell Structure and Function • Guided Reading and Study

Looking Inside Cells (continued)

10. In the table below, describe the function of each organelle in the cytoplasm.

Organelles in Cytoplasm	
Organelle	Function
Mitochondria	Convert energy in food molecules to energy the cell can use to carry out its functions.
Endoplasmic reticulum	Carry proteins and other materials from one part of the cell to another.
Ribosomes	Produce proteins
Golgi bodies	Receive proteins + other materials from the E.R., package them and distribute them.
Chloroplasts	Capture energy from sunlight and use it to produce food for the cell.
Vacuoles	Store food, water and waste.
Lysosomes	Contain chemicals that break down food, old cell parts and the release substances so they can be used again.

Specialized Cells

11. The structure of each kind of body cell is suited to its function.

Bacterial Cells

12. Circle the letter of each sentence that is true about bacterial cells.
- Bacterial cells are larger than plant or animal cells.
 - ☒ Bacterial cells have a cell wall and a cell membrane.
 - ☒ Bacterial cells do not have a nucleus.
 - Bacterial cells do not have genetic material.

cells form tissues that form organs that are part of organ systems that make up an organism.

Looking Inside Cells Worksheet Lesson 2 Answers

Warner, Werner



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Three Key Lessons in Leadership Dave Anderson Corporation,1997 **Looking in Classrooms** Thomas L. Good,Jere E. Brophy,1991 This classic book provides a balanced research based overview of the most effective teaching methods and guides teachers to select appropriate methods by intelligent observation in their classrooms Widely considered to be the most comprehensive and authoritative source available on effective teaching Looking in Classrooms synthesizes the knowledge base on student motivation classroom teaching teacher expectations and adapting instruction for individual learners This book represents state of the art research reviews in several areas including student motivation classroom instruction and student learning classroom management and adapting instruction to the needs of individual students K 12 educators administrators and superintendents *Whizkids Spreadsheets Ii' 2002 Millennium Ed. , New National Framework Mathematics* M. J. Tipler,2004 New National Framework Mathematics features extensive teacher support materials which include dedicated resources to support each Core and Plus Book The 9 Core Teacher Resource Pack contains a wealth of resources to support and extend the work covered in the 9 Core pupil book and Teacher Planning Pack

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web objective to apply stoichiometric principles to calculate the appropriate chemical equation for the reaction between iron
metal and a solution of copper i sulfate introduction in the form of finely split powder copper metal is produced as a
byproduct of the process as iron ions the iron metal dissolves

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web purpose to enhance the understanding of stoichiometry a reaction between iron and copper ii sulfate solution will be
conducted this will help you to differentiate limiting and excess reactant in a chemical reaction finally the theoretical and
percent yield of this reaction will be calculated

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web purpose in this experiment you will observe a single replacement reaction you will also use stoichiometry to predict what
the theoretical yield of product is and calculate a percent yield reaction iron copper ii sulfate copper iron ii sulfate pre lab
questions write the balanced equation

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research was on the irreversible oxidation of trimethylamine by copper ii chloride which takes place at temperatures of about 75 and higher the coordination of trimethylamine by copper ii

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web october 3 2017 abstract the techniques used in this lab are quantitative transfer and vacuum filtration with the reaction of 8 001 grams of copper ii sulfate CuSO_4 and 2 0153 grams of iron powder Fe the goal of this experiment was to determine the product of copper ii sulfate with iron

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web procedure 1 find the mass of a clean dry 100 ml beaker and record 2 measure out 8 00 grams of copper ii sulfate onto a piece of filter paper taking the filter paper s mass into account add these to the beaker 3 measure 50 0 ml of water in the cylinder and add it to the beaker 4

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web stoichiometry lab the reaction of iron with copper ii sulfate the study of stoichiometry deals with the calculation of quantities in a chemical reaction how much product will be produced how much reactant do you need to make that much product

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web description great lab for use during stoichiometry unit in chemistry reinforces ideas of mole ratios and stoichiometry calculations by measuring the mass of iron that reacts and the mass of copper metal produced students can calculate the ratio of moles of reactant to moles of product this experimental mole ratio from lab can be compared

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web in this experiment we will use stoichiometric principles to deduce the appropriate equation for the reaction between metallic iron and a solution of copper ii sulfate this reaction produces metallic copper which is seen precipitating as a finely divided red powder

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web in this lab we will use experimental evidence and the principles of stoichiometry to determine which ion is formed in a single replacement reaction between iron and copper ii sulfate pre lab assignment in preparation for the lab activity we will investigate another single replacement reaction between copper and silver nitrate

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web through this experiment we sought to use stoichiometric principles which derive from stoichiometry the study of substance quantities in ratio or proportion in a chemical reaction in this lab experiment the principles help distinguish the equation between iron powder and aqueous copper ii sulfate in the iron copper ii sulfate redox reaction

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web this experiment is focused on a reaction between iron metal and a solution of copper ii sulfate we are able to figure out the chemical equation between the two by using stoichiometric principles the iron displaces the copper ii sulfate ions the copper ii is then reduced and gains electrons

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web stoichiometry can be defined as the relationship between the relative quantities of substances taking part in a reaction or forming a compound stoichiometric principles can be used to deduce the appropriate chemical equation for a reaction in our experiment we used solid iron in the form of powder was added to a solution of copper ii sulfate

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web stoichiometric principles can be used to determine the chemical equation between iron powder and aqueous copper ii sulfate in the iron copper ii sulfate redox reaction the iron ions dissolve in the copper ii sulfate solution to produce copper ii ions

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web under the finalise of the experiment it was found that iron was the limiting reactant and iron ii sulfate was a product of the reaction between 7 0002g away copper ii sulfate and 2 0101g iron metal which produced 2 3037g of copper product with a percent yield of 100 72 introduction

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web stoichiometry lab report the reaction of iron with copper ii sulfate period 5 in this experiment i observed the reaction of iron with copper ii sulfate i obtained the results to this experiment through adding heated CuSO_4 solution to the iron powder stirring the solution adding water then removing water adding acetone heating the acetone on low

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