

Lewis Structure Worksheet #2**Key**

Draw Lewis Structures for each of the following molecules or ions:

1. Br_2 14 $\text{:}\ddot{\text{Br}}-\ddot{\text{Br}}\text{:}$	2. BH_3 6 $\begin{array}{c} \text{H} - \text{B} - \text{H} \\ \\ \text{H} \end{array}$
3. O_2 12 $\text{:}\ddot{\text{O}}=\ddot{\text{O}}\text{:}$	4. PS_3 24 $\begin{array}{c} \text{:}\ddot{\text{S}}=\text{P}-\ddot{\text{S}}\text{:} \\ \\ \text{:}\ddot{\text{S}}\text{:} \end{array} + \text{resonance}$
5. H_2CS 12 $\begin{array}{c} \text{H} \\ \diagdown \\ \text{C} = \ddot{\text{S}}\text{:} \\ \diagup \\ \text{H} \end{array}$	6. CHCl_3 26 $\begin{array}{c} \text{:}\ddot{\text{Cl}}\text{:} \\ \\ \text{H} - \text{C} - \ddot{\text{Cl}}\text{:} \\ \\ \text{:}\ddot{\text{Cl}}\text{:} \end{array}$
7. SF_2 20 $\text{:}\ddot{\text{F}}-\ddot{\text{S}}-\ddot{\text{F}}\text{:}$	8. SHF 14 $\text{H}-\ddot{\text{S}}-\ddot{\text{F}}\text{:}$
9. N_2F_4 38 $\begin{array}{c} \text{:}\ddot{\text{F}}\text{:} \quad \text{:}\ddot{\text{F}}\text{:} \\ \quad \\ \text{:}\ddot{\text{N}}-\text{N}\text{:} \\ \quad \\ \text{:}\ddot{\text{F}}\text{:} \quad \text{:}\ddot{\text{F}}\text{:} \end{array}$	10. XeO_4 32 $\begin{array}{c} \text{:}\ddot{\text{O}}\text{:} \quad \text{:}\ddot{\text{O}}\text{:} \\ \diagdown \quad \diagup \\ \text{Xe} \\ \diagup \quad \diagdown \\ \text{:}\ddot{\text{O}}\text{:} \quad \text{:}\ddot{\text{O}}\text{:} \end{array}$
11. N_2O_3 (dinitrogen trioxide)* 28e $\text{:}\ddot{\text{O}}=\ddot{\text{N}}-\text{N}=\ddot{\text{O}}\text{:} + \text{resonance}$	12. NO_2 (nitrogen dioxide)* 17e $\text{:}\ddot{\text{O}}=\ddot{\text{N}}-\ddot{\text{O}}\text{:} + \text{resonance}$
13. N_2O_4 (dinitrogen tetraoxide)* 34 $\begin{array}{c} \text{:}\ddot{\text{O}}\text{:} \quad \text{:}\ddot{\text{O}}\text{:} \\ \quad \\ \text{:}\ddot{\text{N}}-\text{N}\text{:} \\ \quad \\ \text{:}\ddot{\text{O}}\text{:} \quad \text{:}\ddot{\text{O}}\text{:} \end{array} + \text{resonance}$	14. N_2O_5 (dinitrogen pentaoxide)* 40e $\begin{array}{c} \text{:}\ddot{\text{O}}\text{:} \quad \text{:}\ddot{\text{O}}\text{:} \quad \text{:}\ddot{\text{O}}\text{:} \\ \quad \quad \\ \text{:}\ddot{\text{N}}-\text{N}\text{:} \\ \quad \\ \text{:}\ddot{\text{O}}\text{:} \quad \text{:}\ddot{\text{O}}\text{:} \end{array} + \text{resonance}$

*Problems 11-14 are more challenging and may have more than one correct structure. Please note that if a molecule has an odd number of valence electrons, at least one atom in the molecule will NOT fulfill the octet rule.

Lewis Structure Worksheet Guided Practice Answers

Ying Liu



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