

Lewis Dot Structures of Atoms and Ions

Why?

The chemical properties of an element are based on the number of electrons in the outer shell of its atoms. We use Lewis dot structures to map these *valence electrons* in order to identify stable electron configurations. Most atoms tend to lose or gain electrons in order to achieve *noble gas stability*. For example, the metal sodium will react violently with the gas chlorine, but once the elements form the compound sodium chloride (table salt) they are very stable due to the stable electron configuration.

Learning Objectives

- Use Lewis dot structures to represent the valence shells of metal and nonmetal atoms and ions.

Success Criteria

- Accurate drawings of Lewis dot structures for atoms and ions.

Resources

- Periodic table including electron configurations
- Reference material containing the names and symbols of elements

New Concept

- The stability of the noble gas electron configuration of 8 electrons

Prerequisites

- Definition of an atom and an atomic ion
- Knowledge of valence electrons
- Knowledge of charges on ions formed by metals and nonmetals
- Definitions of ionization energy and electronegativity
- Knowledge that positive and negative charges strongly attract each other.

Task

Watch the video clip at:

<http://chemed.chem.wisc.edu/JCESoft/CCA/CCA0/Movies/NACL1.html>

Key Questions

1. What is the product of sodium metal reacting with chlorine gas?
2. Is energy released or absorbed in the reaction?

Lewis Dot Structure Of Atoms And Ions Pogil Answer

JS Bruner



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