

Test for Anions :

| Anion | Test | Result |
|---|---|--|
| Carbonate (CO_3^{2-}) | Add dilute acid (e.g. dilute hydrochloric acid) | Effervescence, CO_2 produced (turns limewater milky) |
| Sulfite (SO_3^{2-}) | Add dilute acid (e.g. dilute hydrochloric acid), warm | SO_2 produced (turns aq. potassium manganate (VII) from purple to colourless) |
| Chloride (Cl^-) [in solution] | Acidify, then add aq. silver nitrate [Ag^+] Acidify, then add aq. lead nitrate [Pb^{2+}] | White ppt. of AgCl $\text{Ag}^+ (\text{aq}) + \text{Cl}^- (\text{aq}) \rightarrow \text{AgCl} (\text{s})$ White ppt. of PbCl_2 |
| Bromide (Br^-) [in solution] | Acidify, then add aq. silver nitrate | Cream ppt. of AgBr |
| Iodide (I^-) [in solution] | Acidify, then add aq. silver nitrate (or lead nitrate) | Yellow ppt. of AgI (or PbI_2) |
| Sulfate (SO_4^{2-}) [in solution] | Acidify, then add aq. barium nitrate (or barium chloride) [Ba^{2+}] | White ppt. of BaSO_4 $\text{Ba}^{2+} + \text{SO}_4^{2-} \rightarrow \text{BaSO}_4$ |
| Nitrate (NO_3^-) | Add aq. sodium hydroxide, then little aluminium (foil) and warm carefully | Ammonia produced (turns damp red litmus paper or universal indicator paper blue) |

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