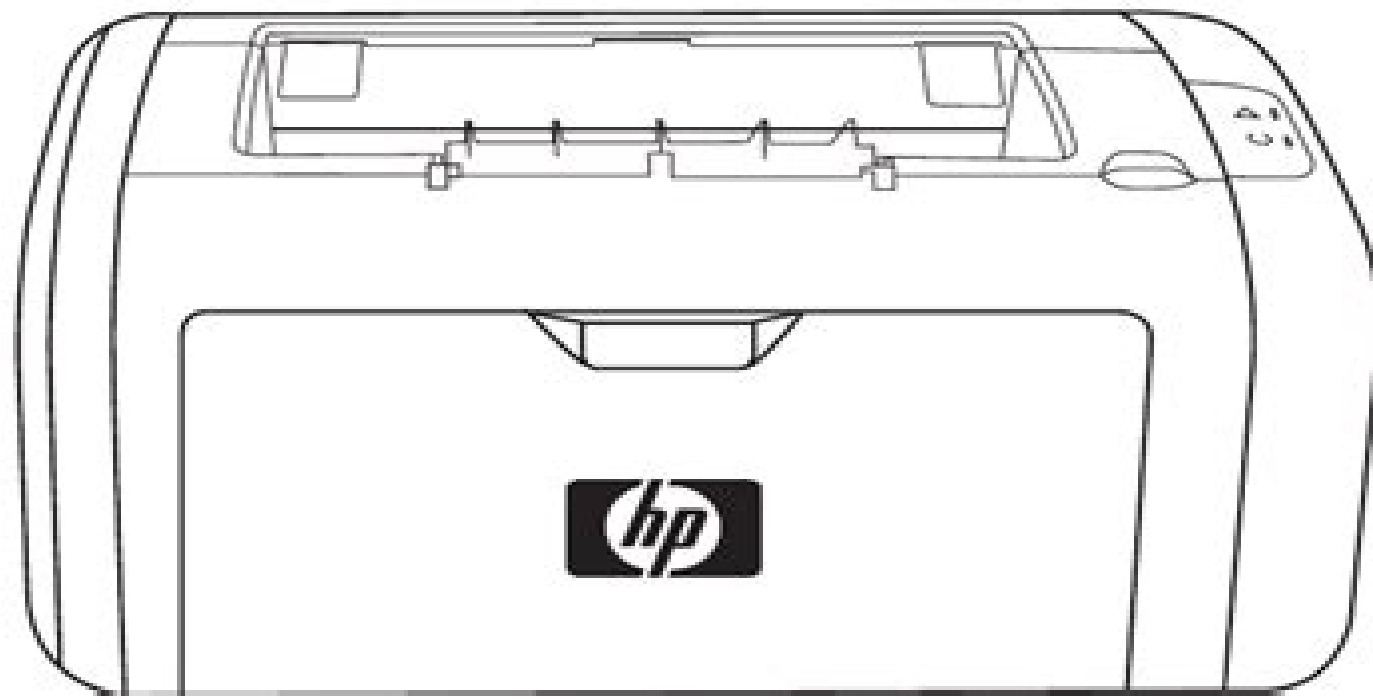




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naoh b  $1.15 \times 10^3$  g lactose c  $82.1$  g fe d  $0.332$  g no  $2.3654 \times 10^{21}$  molecules of c  $3$  h  $8$  o  $3.4279$  grams of nh  $3.5197 \times 10^{22}$  grams of sn

*chapter 10 chemical quantities scarsdale public schools* - Jan 13 2023

web chapter 10 chemical quantities basics the basic unit that is used to determine the amount of a chemical substance is called a mole a mole mol of a substance is equivalent to  $6.02 \times 10^{23}$  particles of that substance

chemistry 12th edition chapter 10 chemical quantities 10 1 - Feb 14 2023

web chemistry 12th edition answers to chapter 10 chemical quantities 10 1 the mole a measurement of matter 10 1 lesson  
check page 315 9 including work step by step written by community members like you