

## Definitions

### Atom -

Particles that make up all substances.

### Element -

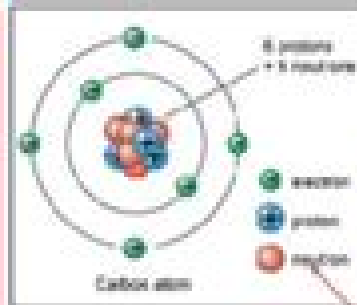
A substance made up of only one kind of atom.

### Compound -

A substance made of different types of atoms joined together.

### Mixture -

A substance consisting of atoms that can be easily separated as they are not joined together.



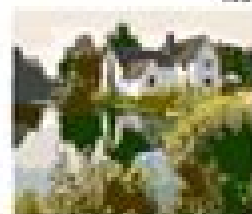
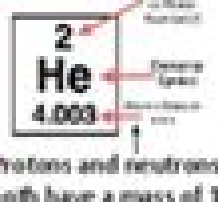
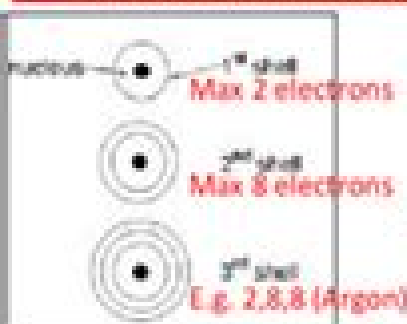
Periodic Table of the Elements

|    |    |    |    |    |    |    |    |    |    |    |    |    |    |    |    |    |    |
|----|----|----|----|----|----|----|----|----|----|----|----|----|----|----|----|----|----|
| 1  | 2  | 3  | 4  | 5  | 6  | 7  | 8  | 9  | 10 | 11 | 12 | 13 | 14 | 15 | 16 | 17 | 18 |
| H  | He | Li | Be | B  | C  | N  | O  | F  | Ne | Na | Mg | Al | Si | P  | S  | Cl | Ar |
| K  | Ca | Sc | Ti | V  | Cr | Mn | Fe | Co | Ni | Cu | Zn | Ga | Ge | As | Se | Br | Kr |
| Rb | Sr | Y  | Zr | Nb | Mo | Tc | Ru | Rh | Pd | Ag | Cd | In | Sn | Sb | Te | I  | Xe |
| Cs | Ba | La | Hf | Ta | W  | Re | Os | Ir | Pt | Au | Hg | Tl | Pb | Bi | Po | At | Rn |
| Fr | Ra | Ac | Th | Pa | U  | Np | Pu | Am | Cm | Bk | Cf | Es | Fm | Md | No | Lr |    |

No charge

Group 0 or 8 are **UNREACTIVE**  
They have a full outer shell of electrons

Periodic table arranges elements by **ATOMIC NUMBER** (proton number/the small one)



Used for building

**LIMESTONE**

Calcium Carbonate **CaCO<sub>3</sub>**

HEAT

(thermal decomposition)

**CaO**

Calcium Oxide

Limewater

The test for CO<sub>2</sub>

**Ca(OH)<sub>2</sub>**

Or Calcium Hydroxide

comes from adding water to CaO

**Carbonates**  
Decompose when heated  
= Oxide + CO<sub>2</sub>  
React with acid  
= Salt + water + CO<sub>2</sub>

## Balanced Equations



C=1

H=4

O=4

C=1

H=4

O=4

A balanced equation has the same number of atoms of each element on both sides

Using Molecules: In this equation, the large number in front of the chemical symbol tells you the molecules needed.

e.g. 1 CH<sub>4</sub> molecule reacts with 2 O<sub>2</sub> molecules

## Covalent Bonding

Non-metals sharing electrons



## Ionic Bonding



**Good:** More jobs  
**Bad:** Destroy habitats and Landscape  
Would you want to live next to one?

## TRANSITION METALS

**Metals - Very useful**  
e.g. Copper wires conduct well  
Most need to be **ALLOYED** to make them harder

Metals which are un-reactive are found in their **NATIVE STATE** e.g. **GOLD**  
More reactive metals are found as **METALS ORES** and need to be **EXTRACTED**

# C1 CHEMISTRY

# Gcse Chemistry C1 March 2013 Past Papers

**A Loxley**



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