



Lewis Structures and Molecular Geometry

Models of Covalent Bonding

Introduction

Molecules have shaped The structure and shape of a modecule influences, its physical properties and affects its abenical behavior as well. Lewis structures and VSEPR theory offer useful models for visualizing the structures of covalent compounds.

Concepts

Waterson electronic

Considers booking:

* Legeria structures.

· WSEPR discorp

Background

Considers, bornis are defined as the net attractive forces between nonnestal atoms that share one, two in three pains of electrons. In general, only the universe electrons, those in the highest energy levels that are furthest away from the madeau, are available for borning. The number of odente electrons influences the number of borning the number of odente electrons will form a consenters abortout for determining the number of an element in the number of an element in the resident Periodic Table reflects its electron configuration. When the representative electrons are arranged in columns from IA to VIIIA have Figure II, the number of salence electrons for an element is exact to its group number. Thus, potassium in Group IA has one call to its group number. Thus, potassium in Group IA has one chlusters electrons, carbon in Group IVA has four electrons.



Figure 1. Numbering of representative elements in Groups IA-VIIIA.

In 1916, C.N. Levets, an American chemist, proposed arranging distaaround the symbols of the elements to represent the valence electrons. Levals electron and specials (see Figure 2) remain the most popular may unplicate the valence electrons that are available for borning.



Figure 2. Lewis electronyclot syndrols for representative elements.

Accountmentation build on the Lewis electron-dost symbols of the elements to show the bonding arrangement of atoms in a molecule and the distribution of all valence electrons. The Lewis structure of a molecule thus shows all of the atoms and how they are connected. A single totaless band between two atoms, corresponding to a pair of electrons, is represented using a dosh (—). Sometimes atoms shore none than one pair of electrons them in order to form stable molecules. Two dashes (corresponding to their pairs of electrons) and three dashes (corresponding to three pairs of electrons) are used to represent double and triple lesses, respectively.

63. N. Lerwis riffered: a simple theory, hased on the known stability of the robble gases (Ple, Ne, etc.), to precifet how many bonds an atom well form and how many atoms of a particular type will come together to form a stable routecale. According to Lerwis, nonmertals may share electrons in order to achieve a valence obell electron "cours." similar to that of the robble gases.

""Elso atomic many conforms to the rade of eight, or the votet rade...
By when ing one or more pains of electronic. The electronic tablesh
are held in common by those atomic may be considered to belong
to the water shell of both atomic."

The recible gases have filled a and provisials with eight electrons. The octavinal measurement that atoms force replectable to achieve this stable, recible gas electron configuration. In counting submodefectrons in order to precise the structure of a constent compounts, we will distinguish between two birds of electrons pains. Broading pains of electrons are shared between atoms and thus "belong" to both atoms in the bond. Newhooding or woodshood pains of electrons are not shared between atoms and pains of electrons are not shared between atoms and are therefore "counters" toward only one of the atoms.

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