

(iii) What does  $x\text{H}_2\text{O}$  represent in the formula  $\text{Al}_2(\text{SO}_4)_3 \cdot x\text{H}_2\text{O}$ ?

Water of crystallisation.

(1)

$\text{H}_2\text{O}$  crystals to constant mass.

has a mass of 5.76g.

value of  $x$ .

+



$$\frac{5.76\text{g}}{18}$$

$$= 0.32\text{mol}$$

$$\frac{0.02}{16}$$

$$= 16$$

# June 2013 Past Paper F321

**M Woodhall**



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