

**Assessment 1****Chapter Test B****Chapter: Chemical Equations and Reactions**

**Part 1** In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

- \_\_\_\_\_ 1. The precipitation of a slightly soluble solid compound from a solution of two ionic reactants is called the formation of a  
a. gas.  
b. precipitate.  
c. combustion reaction.  
d. halogen.
- \_\_\_\_\_ 2. To balance a chemical equation, it is permissible to adjust the  
a. coefficients.  
b. subscripts.  
c. formulae of the products.  
d. number of products.
- \_\_\_\_\_ 3. In a chemical equation, the symbol (aq) indicates that the substance is  
a. melted.  
b. dissolved in water.  
c. an acid.  
d. insoluble.
- \_\_\_\_\_ 4. The tendency for a replacement reaction to occur depends on the  
a. relative tendency of the two elements to lose electrons.  
b. temperature difference.  
c. relative electronegativity values.  
d. relative tendency of the two elements to gain electrons.
- \_\_\_\_\_ 5. The coefficients in a chemical equation  
a. indicate the number of moles of each substance that react.  
b. show the number of grams of each substance that react.  
c. are the molar masses of the substances.  
d. show the relative electronegativity values.
- \_\_\_\_\_ 6. If metal X is more reactive than metal Y in the activity series, then  
a. X will replace ions of Y in solution.  
b. Y will replace ions of X in solution.  
c. Y will displace cations only indirectly.  
d. X will react with cold water.

# Holt Chemistry Study Guide Reaction Rate Answers

**Daniel F McAuley**



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