

1. Consider the following reaction: $N_2(g) + 3H_2(g) \longrightarrow 2NH_3(g)$, if the rate of the loss of hydrogen gas is $0.09 \text{ mol/l}\cdot\text{s}$, what is the rate of production of ammonia?

2. The rate law of a given reaction is $r = kC^2$, with $k = 2.64 \times 10^{-3} \text{ M}^{-1} \text{ min}^{-1}$, If the initial concentration is 0.025 M , what is the initial rate?

3. the reaction $2NO(g) + Cl_2(g) \longrightarrow 2NOCl(g)$, with rate law, $r = kC_{NO}^2C_{Cl_2}$ is _____ order with respect to NO, _____ order with respect to Cl_2 and _____ order overall.

4. The rate of a particular reaction is quadrupled when the temperature was increased from 55°C to 60°C . What is activation energy (in J/mol) for this process.

5. A common rule of thumb is that near room temperature, the rate of a reaction will double with each 10°C rise in temperature. Estimate the activation energy (in J/mol) needed for this rule to hold true.

6. The mechanism of a reaction is shown below:



- (A) What is the overall reaction?
(B) Which compounds are intermediates?
(C) Predict the rate law based on this mechanism.
(D) What is the overall order of the reaction?

7. consider the following reaction, $N_2H_4 + 2H_2O_2 \longrightarrow N_2 + 4H_2O$. In 55 seconds, 0.165 mol of H_2O_2 is consumed. The rate of production of N_2 is

- (A) $1.5 \times 10^{-3} \text{ mol}$ in 55 seconds
(B) 0.33 mol/s
(C) $1.5 \times 10^{-3} \text{ mol/s}$
(D) 0.006 mol/s

8. The dimensions of rate constant for reaction $3A \longrightarrow B$ are $(\text{l/gm mole})/\text{min}$. Therefore the reaction order is

- (A) 0
(B) 1
(C) 2
(D) 3

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