

4. Which of the following substances would have hydrogen bonding as one of its IMF's?



5. Which molecular substance from the following list would have the highest boiling point?

- A.  $\text{Ne}$
- B.  $\text{F}_2$
- C.  $\text{H}_2$
- D.  $\text{Cl}_2$

6. When the actual gas pressure is less than the pressure predicted by the ideal gas law, the explanation lies in the fact that the ideal gas law does NOT include a factor for molecular

- A. volume
- B. mass
- C. velocity
- D. attraction
- E. shape

7. 25.0 L of hydrogen gas is collected over water at  $40^\circ\text{C}$ . The total pressure of the sample is 750 mmHg. At  $40^\circ\text{C}$ , the vapor pressure of water is 55 mmHg. What is the mass of the hydrogen gas?

- A. 1.88 g
- B. 1.55 g
- C. 0.896 g
- D. 0.967 g
- E. 1.12 g

8. A 2.08 L sample of nitrogen gas at  $27^\circ\text{C}$  and 680 mmHg is heated until it occupies a volume of 3.80 L. If the pressure remains unchanged, the final temperature of the gas is

- A.  $48^\circ\text{C}$
- B.  $120^\circ\text{C}$
- C.  $477^\circ\text{C}$
- D.  $677^\circ\text{C}$
- E.  $990^\circ\text{C}$

9. PI (d) bonding occurs in each of the following EXCEPT

- A.  $\text{C}_2\text{H}_2$
- B.  $\text{C}_2\text{H}_4$
- C.  $\text{C}_2\text{H}_6$
- D.  $\text{C}_2\text{H}_2$
- E.  $\text{C}_2\text{H}_4$

10. Determine the shape and polarity of the  $\text{PF}_3$  molecule

- A. trigonal pyramidal, polar
- B. trigonal pyramidal, nonpolar
- C. trigonal planar, polar
- D. trigonal planar, nonpolar

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**KJ Lindholm-Leary**



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