

# Unit 1 Test Study Guide: Geometry Basics

**What is  
the format?**

## Part I - Vocabulary

There will be 15 vocabulary words taken from your dictionary.

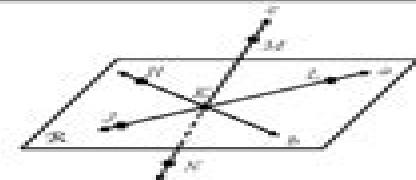
## Part II - Short Answer

There will be 25 short answer questions based all topics covered this unit.

### Topic #1: Points, Lines, & Planes

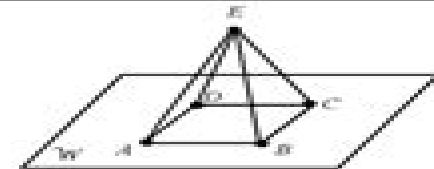
Use the diagram to the right to answer questions 1-4.

1. Name two points collinear to point  $K$ . \_\_\_\_\_
2. Give another name for line  $b$ . \_\_\_\_\_
3. Name the intersection of line  $c$  and plane  $R$ . \_\_\_\_\_
4. Name a point non-coplanar to plane  $R$ . \_\_\_\_\_



Use the diagram to the right to answer questions 5-8.

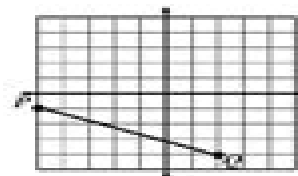
5. How many planes are shown in the figure? \_\_\_\_\_
6. Give another name for plane  $W$ . \_\_\_\_\_
7. Name the intersection of plane  $ADE$  and plane  $W$ . \_\_\_\_\_
8. Name a point non-collinear to points  $A$  and  $B$ . \_\_\_\_\_



### Topic #2: Distance & Midpoint

9. Find  $ST$  if  $S(-3, 10)$  and  $T(-2, 3)$ .

10. Given the graph below, find  $PQ$ .



11. Find the coordinates of the midpoint of  $HK$  if  $H(-1, 2)$  and  $K(-7, -4)$ .

12. Find the coordinates of  $Z$  if  $Y$  is the midpoint of  $XZ$ ,  $X(-10, 9)$  and  $Y(-4, 8)$ .

# Geometry Chapter 1 Unit Test Answers

**Kaplan Nursing**



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label each as oxidation or reduction 5 al 6 s 7 o 2 8 ba2 9 n3 10 br2 11 p 12

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web 1 define each remember oil rig oxidation is loss of e reduction is gain of e a oxidation b reduction c oxidizing agent d reducing agent 2 write half reactions for each of the following atoms or ions label each as oxidation or reduction use the activity series chart a al b ba2 c br 2 d ca e ga3 f h 2 g h

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web during oxidation and reduction the oxidation numbers of atoms change if an atom is oxidized its oxidation number increases ie it becomes more ve or less ve if an atom is reduced its oxidation number decreases ie it becomes less ve or more ve

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web you should try to answer the questions without referring to your textbook if you get stuck try asking another group for help 1 identify the oxidation state of each element in the following a na 2 so 4 b mno 4 c kclo 3 d ch 3 oh 2 identify the oxidation state of nitrogen in the following a no 3 b no 2 c no 2

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web an oxidation reduction redox reaction is one in which one or more atoms change oxidation numbers oxidation occurs

when an atom's oxidation state becomes more positive indicating that electrons have been lost reduction occurs when an atom's oxidation state becomes less positive indicating that electrons have been gained

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web describe redox reactions in terms of electron transfer recall that oxidation involves loss of electrons and reduction involves gain of electrons represent simple redox reactions using half equations identify oxidising and reducing agents for simple redox reactions understand the reactivity of metals and metal displacement reaction

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web what is an oxidation reduction or redox reaction this is a reaction in which electrons are transferred between reactants what do the following terms mean oxidation oxidation is a loss of electrons reduction reduction is a gain of electrons what is a helpful way to remember this

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web write the oxidation and reduction reactions for each redox reaction the first one is done for you 23  $\text{Fe} + \text{Co} \rightarrow \text{Fe}^{2+} + \text{Co}^{2+}$  oxidation  $\text{Co} \rightarrow \text{Co}^{2+} + 2\text{e}^-$  reduction  $\text{Fe}^{2+} + 2\text{e}^- \rightarrow \text{Fe}$  24 3  $\text{Ag} + \text{Ni} \rightarrow \text{Ag} + \text{Ni}^{2+}$  oxidation reduction 25  $\text{Cu} + \text{Pb} \rightarrow \text{Cu}^{2+} + \text{Pb}$  oxidation reduction 26  $\text{O}_2 + 2\text{H}_2 \rightarrow 2\text{H}_2\text{O}$

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web in each of the following redox reactions indicate the elements that undergo changes in oxidation number and the magnitude of change then identify the element oxidized and reduced as well as the oxidizing agent and reducing agent a  $\text{I}_2 + 5\text{O}^{2-} \rightarrow 2\text{IO}_3^- + 5\text{O}^{2-}$  b  $\text{Co} + \text{I}_2 + 5\text{O}^{2-} \rightarrow \text{Co}^{2+} + 2\text{IO}_3^- + 5\text{O}^{2-}$

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web chemical reactions in which there is a transfer of electrons from one substance to another are known as oxidation reduction reactions or redox reactions in this unit you will examine the oxidation reduction process and use the oxidation state and oxidation number concepts to identify redox reactions and to keep track of electrons

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web jun 28 2010 answers i determine what is oxidized and what in each reaction identify the oxidizing agent and the reducing agent also  $2\text{Sr} + \text{O}_2 \rightarrow 2\text{SrO}$   $\text{Sr}^0$  to  $\text{Sr}^{+2}$  oxidized reducing agent  $\text{O}^0$  to  $\text{O}^{-2}$  reduced ox ag  $2\text{Li} + \text{S} \rightarrow \text{Li}_2\text{S}$   $\text{Li}^0$  to  $\text{Li}^{+1}$  oxidized red ag  $\text{S}^0$  to  $\text{S}^{-2}$  reduced ox ag 3  $2\text{Cs} + \text{Br}_2$

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