



Formula of an Ionic Compound

Balancing Charges on Ions

Introduction

Atoms of different elements conshine with one another to form composeds. The empirical formula of an ionic compound indicates the kinds of atoms that are present in the compound as well as the relative number (ratio) of each kind of atom. Let's investigate how the formula of an ionic compound can be determined experimentally.

Concepts

• Тогейс спотерхоштеся

· Engirical formula

Profysikowińci i ocea-

Precipitation reaction:

Background

An intercomposited is comprised of intermations or groups of atoms that have a positive or negative charge. Oppositely charged iron arrange themselves into an extended, three conveniental structure called a crystal lattice. The net attractive forces among oppositely charged irons in the crystal structure are called torric bonds. Although composed of charged irons, ionic composeds are electrically neutral. The ratio of oppositely charged iron in the crystal structure is such that the positive charge contributed by the cational around to or balanced by the negative charge contributed by the anison. There is no net or overall charge in an ionic composers.

The empirical fermion of an ionic conspound includes the smalless, whole-mander national each type of ion in the crystal structure and is called a formula unit. For example, magnesium otherwise has the engine-cal formula MgCI₂. Magnesium cutions (MgI) and obtains anions (CI) considere in a 1:2 each to form the MgCI₂ formula unit. The overall tharpe on ionic compounds is always seen.

Some inner consists of a charged group of covalently borsied atories. Such tions are called polyatoxic inter. An example is the nitrate ion $(NO_{\alpha})_{\alpha}$ which contains one mitrogen atore and three oxygen atores, and has an everall charge of -1. In calcium other tie, addison $(Ca^{\alpha})_{\alpha}$ items condition with retrate ions in a 1-2 ratio in order to halone the positive and negative charges. The empirical formula for calcium nitrate is $Ca(NO_{\alpha})_{\alpha}$. Parentheses are used around the nitrate ion to show that the subscript $Ca(NO_{\alpha})_{\alpha}$ pertains to the nitrate ion as a whole.

Many torsic conspounds can be prepared in the lab using projectives reserves. When substitute of two torsic compounds are consistent, the insis may reserve to form a new torsic compound that is insulable in water. An example of this type of reaction is the formation of solid barisms sulfate when because of solid prisate and solid because the consistent of solid perisate solution (Equation 1.4). In Equation 1b, only the ions that form the pre-cipitate are represented. This makes it easier to recognize what bappens in the precipitation resultion.

$$\mathbb{R}s^{t}(sc) + SO_{s}^{s}(sc) \rightarrow \mathbb{R}sSO_{s}(s)$$
 Equation 18

According to the balanced equation for this reaction, barians time (Ba^2) , combine with subface irons (SO_a^2) in a 1:1 ratio to form became subface $(BaSO_a)$. This ratio can be observed experimentally in the lab by mining $BaCI_b$ tact and Na_aSO_a tact substitute containing exact assumes (concentrations) of bettern and subface tone, respectively. The maximum around of precipitate will be obtained when exact substances (a.1: train) of the two solutions are condition. A similar approach can also be used to determine the formula of an unknown tonic companies.

Experiment Overview

The purpose of this experiment to to determine the empirical formula of an unforcer investigation. Two solutions consisting equal amounts borners around in the new action of the own relations will be conditionation will be held unready while the volume of the two solutions will be held unready, while the volume nation of the ready-ried. The amount of precipitate obtained in each reaction will be necessarily while the volume and reaction will be necessarily and placed and placed.

Flinn Chemtopic Labs Ionic Compound Answers

Tao Wei

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