

**M** A buffer is made using the conjugate acid-base pair acetic acid and acetate ions. Write chemical equations showing the reactions that take place when  $H^+$  and when  $OH^-$  are added to the buffer.

## Materials

### Chemical Analysis

Acetic acid,  $\text{HC}_2\text{H}_3\text{O}_2$ , 0.1 M, 20 ml.

Ammonia, NH<sub>3</sub>, 0.1 M, 20 ml.Ammonium chloride,  $\text{NH}_4\text{Cl}$ , 0.1 M, 20 ml.

Hydrochloric acid (HCl, 0.1 M, 5 ml)

Phenylhydrazine, 1.0 g, 1 ml

Sodium acetate,  $\text{NaCH}_3\text{CO}_2$ , 0.1 M, 20 ml.

**Abstract**

Halderson, C. A. (ed.)

Features: 100- and 400-

Abstract—Cancer reports. 3

Buret, ring stand, and buret clamp

Edinburgh 1960, 150 ml. 2

Sodium chloride, NaCl, 0.1 M, 40 ml.

Sodium Hydroxide, NaOH, 0.1 M, 5 ml.

Sodium hydroxide, NaOH, 0.2 M, 30 ml.

Unknown solid acid anhydride, 2.0 g

Water, distilled or deionized

$$(0.057M)(.100L) = 0.0057\text{mol}$$

$$[C_2H_3O_2^-] = \frac{K_a [HC_2H_3O_2]}{[H^+]} = \frac{(1.8 \times 10^{-5})(0.10)}{3.16 \times 10^{-5}}$$

Wash bottle with distilled or deionized water.

# Flinn Buffer Solutions Lab Answer Key

**Rachel S Tattersall**



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