

Chemistry Unit 7 Worksheet 4

Samples of Every Kind of Problem

On a separate sheet of paper, write a complete solution to each of the problems below. Follow the procedures outlined in class. Be sure to circle your final answer.

- Calculate the number of moles of potassium chlorate (KClO_3) that must decompose to produce potassium chloride (KCl) and 1.5 moles of oxygen gas.
- In a single displacement reaction, magnesium metal reacts with hydrochloric acid to produce magnesium chloride and hydrogen gas. How many moles of hydrochloric acid are needed to completely react with 1.43 g of magnesium?
- Ethanol ($\text{C}_2\text{H}_5\text{OH}$) reacts with oxygen gas to produce carbon dioxide gas and water vapor. What mass of oxygen gas is required to react with 2.88 moles of ethanol?
- Determine the mass of sodium chloride produced when 0.75 g of sodium (Na) reacts with sodium hydroxide according to the following equation:
 $\text{NaClO}_3 + 2 \text{NaOH} \rightarrow \text{NaClO}_2 + 2 \text{NaOH}$
- In the copper-silver nitrate lab, copper metal and silver nitrate solution reacted to produce silver metal and copper(II) nitrate in solution. A student placed a copper wire with a mass of 2.88 g in the reaction test tube. The silver-silver nitrate solution contained 0.40 g of silver nitrate. She obtained 0.87 g of silver metal. Calculate the percent yield of silver.
- When hydrochloric acid (HCl) is added to sodium hydroxide solution, the products are water, sodium chloride and sodium chloride gas. What is the per cent yield of 4.88 g of Cl_2 gas collected when 20.0 g of sodium hydroxide solution reacts with excess HCl?
- Phosphorus and bromine react (chemically) together to form phosphorus tribromide. If 0.4 g of phosphorus and 10 g of bromine react, how many grams of PBr_3 could be produced?
- Iron sulfate and oxygen gas react to form iron oxide and sulfur dioxide. Determine the amount of SO_2 that should be produced in a reaction between 40.0 g of FeSO_4 and 3.0 g of oxygen. What is the percent of the experiment?

$$1) 1.5 \text{ moles of } \text{KClO}_3 \quad 2) 1.43 \text{ moles of } \text{HCl} \quad 3) 2.88 \text{ g } \text{C}_2\text{H}_5\text{OH} \quad 4) 1.43 \text{ g of } \text{NaClO}_3$$

$$5) 100\% \quad 6) 100\% \quad 7) 100\% \text{ PBr}_3 \quad 8) 100\% \text{ g of } \text{SO}_2 \quad 9) 100\% \text{ g of } \text{FeSO}_4$$

Modeling Chemistry Unit 7 Ws 4

Karin Nielsen-Saines



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