

Chemistry -- Unit 6 Worksheet 1

After having observed molecular models which show 18.098 compounds, you discovered that certain particles are known to form positively charged ions, but cannot, within their natural particles, form negatively charged ions. (Lithium, Sodium, Potassium, Calcium, Barium, Strontium, Magnesium, Beryllium, and Radium). After having observed the 18.098 compounds, you also observed that certain particles are known to form negatively charged ions, but cannot, within their natural particles, form positively charged ions. (Fluorine, Chlorine, Bromine, Iodine, and Astatine). After having observed the particles that form the ions, you discovered that certain particles are known to form positively charged ions, but cannot, within their natural particles, form negatively charged ions. (Lithium, Sodium, Potassium, Calcium, Barium, Strontium, Magnesium, Beryllium, and Radium).

1. After having observed the particles that form the ions, you discovered that certain particles are known to form positively charged ions, but cannot, within their natural particles, form negatively charged ions. (Lithium, Sodium, Potassium, Calcium, Barium, Strontium, Magnesium, Beryllium, and Radium).

18.098 molecular particles	1. Lithium 1. Sodium	1. Potassium 1. Calcium	1. Barium 1. Strontium	1. Radium 1. Magnesium
	Li^+	Na^+	K^+	Ca^{2+}

particles
molecular



18.098 molecular particles	1. Barium 1. Strontium	1. Magnesium 1. Calcium	1. Radium 1. Magnesium
	Ba^{2+}	Mg^{2+}	Ca^{2+}

particles
molecular



18.098 molecular particles	1. Fluorine 1. Chlorine	1. Bromine 1. Iodine	1. Astatine 1. Tellurium	1. Selenium 1. Tellurium
	F^-	Br^-	I^-	At^-

particles
molecular



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Sabine Zange



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