

$$\begin{aligned}
 K' &= \frac{[\text{H}_2\text{O}_2]}{[\text{NO}_2]^2} \\
 &= \frac{(9.3 \times 10^{-3})}{[\text{NO}_2]^2} \\
 &= \frac{(9.3 \times 10^{-3})}{(66.5 \times 10^{-3} \text{ mol/L})^2} \\
 &= 2.20 \\
 &= 2.2 \times 10^2
 \end{aligned}$$

45. (a) Students should expect it to be a base.



- (c) Candy contains sucrose, which will remove hydroxyapatite from teeth by reacting with it at low pH to give $\text{Ca}^{2+}_{(\text{aq})}$ and $\text{H}_2\text{PO}_4^{-}_{(\text{aq})}$ that are in soluble form.

46.

	$\text{NH}_3 \cdot (\text{CH}_3)_2\text{NH}_2$	$\text{NH}_3^+ \cdot (\text{CH}_3)_2\text{NH}_2^+$	OH^-
initial conc.	0.10	0	0
change	-0.0021	+0.0021	+0.0021
equilibrium	0.0979	0.0021	0.0021

$$K_b = \frac{(0.0021)^2}{0.0979} = 4.5 \times 10^{-5}$$

47. The conductivity of 0.1 mol/L HCl is higher than for 0.1 mol/L CH_3COOH because the HCl is completely ionized whereas the CH_3COOH is only partly ionized. For solutions of these two acids at a concentration of 1×10^{-7} mol/L, the HCl is still ionized completely and the CH_3COOH again is only partly ionized. However at such a low concentration the conductivity of the HCl is the same as for water. There will be no discernable difference in the conductivities at this low concentration.

48. One drop of concentrated acid will have a high concentration of H_3O^+ ions. (E.g., For HCl $[\text{H}_3\text{O}^+] =$ about 12.5 mol/L, for H_2SO_4 $[\text{H}_3\text{O}^+] =$ about 35 mol/L. When one drop of acid is added to a bucket of water, the H_3O^+ ion concentration is diluted by a very large factor.

49. (a) A triprotic acid is an acid that has three protons that may be dissociated, each with its own K_a value.

- (b) Stepwise Equation 1:



- Stepwise Equation 2:



- Stepwise Equation 3:



- (c) $\text{H}_2\text{PO}_4^{-}_{(\text{aq})}$ acts as an acid in equation 2 above. In the equation below, it acts as a base.



- (d) Students may say that $\text{H}_3\text{PO}_{4(\text{aq})}$ is the stronger acid because it is higher in the list of relative strengths of acids and bases. Some students may suggest that $\text{H}_3\text{PO}_{4(\text{aq})}$ is a stronger acid because its K_a value is larger than that of $\text{H}_2\text{PO}_4^{-}_{(\text{aq})}$, which is a "stronger" answer.

50. Water could have a pH greater than 7.0 at a temperature different from 25°C.

51. (a) Students might suggest either of these equations:



- (b) The litmus test indicates that the solution is basic, so the solution contains a considerable $[\text{OH}^-]$. Neither of these equations, written according to the original Arrhenius theory, is consistent with the litmus test. Equation two above even implies a red litmus result! Accordingly, the litmus test contradicts the Arrhenius theory equations.

- (c) A student could have two frames of thought.

Perspective I:



or, without the spectator, Na^+ :



Perspective II:



or, without the spectator, Na^+ :



Note that, if a student writes equations according to Perspective I, the student still has a set of equations that contradict the litmus test. But if the student then realizes that the hydrogen carbonate ion is amphoteric, another valid set of equations according to Perspective II may be written. Given the litmus test, the student should then analyze and evaluate the equations to state that Perspective II is not only correct according to modernized Arrhenius, but also consistent with the litmus test. Some students will have realized this out of part (b), and will present only Perspective II.

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**Gordon James Thorogood, Sarah C.
Finkeldei, Maik Kurt Lang, David
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Handbook on the Physics and Chemistry of the Actinides, 1984 *Indian Journal of Chemistry*, 1986 **Ordered and Disordered Cubic Systems: Pyrochlore to Fluorite, Now and the Horizon** Gordon James Thorogood, Sarah C. Finkeldei, Maik Kurt Lang, David Simeone, 2022-02-18 Journal of the American Chemical Society American Chemical Society, 1984-05 Features definitive articles and communications as well as book and software reviews covering all areas of chemistry

Hydrothermal Mineral Deposits Franco Pirajno, 2012-12-06 This book is intended primarily for exploration geologists and post graduate students attending specialist courses in mineral exploration Exploration geologists are engaged not only in the search for new mineral deposits but also in the extension and re assessment of existing ones To succeed in these tasks the exploration geologist is required to be a generalist of the Earth sciences rather than a specialist The exploration geologist needs to be familiar with most aspects of the geology of ore deposits and detailed knowledge as well as experience play an all important role in the successful exploration for mineral commodities In order to achieve this it is essential that the exploration geologist be up to date with the latest developments in the evolution of concepts and ideas in the Earth sciences This is no easy task as thousands of publications appear every year in an ever increasing number of journals periodicals and books For this reason it is also difficult at times to locate appropriate references on a particular mineral deposit type although this problem is alleviated by the existence of large bibliographic data bases of geological records abstracts and papers on computers During my teaching to explorationists and indeed during my years of work as an explorationist the necessity of having a text dealing with the fundamental aspects of hydrothermal mineral deposits has always been compelling Metallic mineral deposits can be categorised into three great families namely 1 magmatic 2 sedimentary and residual 3 hydrothermal

Nuclear Science Abstracts, 1974 Los Alamos Science, 2000 **Index Medicus**, 2001 Vols for 1963 include as pt 2 of the Jan issue Medical subject headings

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Synchrotron-Based Techniques in Soils and Sediments Balwant Singh, Markus Grafe, 2010-02-18 Over the past 20 years synchrotron based research applications have provided important insight into the geochemical cycling of ions and the chemical and crystallographic properties of minerals in soils and sediments Of particular significance is the understanding of local coordination environments with the use of X ray absorption spectroscopy The high flux and brightness of the X ray beams have allowed researchers to work at environmentally relevant concentrations The use of focusing mirrors and apertures which allow for mapping and trace particle surfaces microbes roots channels and elements at the micron and at a nano meter scale in 2 and 3D have also been a great enhancement to science This book provides the most up to date information on synchrotron based research applications in the field of soil sediment and earth sciences Invited authors provide chapters on a wide range of research topics including multiphase flow and transport processes physical aspects rhizosphere and microbial life biological aspects and dynamics of C N S P and heavy metals and metalloids chemical aspects

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