

Chemistry – Unit 6 Worksheet 1

We have observed evidence that atoms of different elements are different, the most noticeable being their consistently different mass. However, when the elements are broken apart by their respective chemical reactions, the elements are broken apart into smaller particles that are consistent in their composition. The most consistent observation is that the mass of the particles that make up the elements is the same. We will now examine the particles that make up the elements and observe their structure and composition. The structure of the particles is the same for all elements.

1. Write the formula and draw the particle diagram for each compound. The table of atomic weights is given.

Atomic weights of elements	16 Hydrogen 1.008	12 Carbon 12.01	16 Oxygen 16.00	35.5 Chlorine 35.5
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particle



Diagram

Atomic weights of elements	16 Hydrogen 1.008	12 Carbon 12.01	16 Oxygen 16.00	35.5 Chlorine 35.5
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particle



Diagram

Atomic weights of elements	16 Hydrogen 1.008	12 Carbon 12.01	35.5 Chlorine 35.5	16 Oxygen 16.00
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particle



Diagram

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R Bogdan



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