

SCHOOLNAME:-

STUDENT'S NAME:-

CLASS:-

SENIOR CLASS TWO (S.S.S. 2)

SUBJECT:-

PHYSICS 1ST - 3RD TERM EXAM

18.
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SECTION A: FIRST TERM QUESTIONS

1. A phenomenon which is used to describe the process by which a substance changes from solid to gas is called (a) evaporation (b) freezing (c) sublimation (d) melting (e) vaporization
2. A phenomenon which is used to describe the process by which a substance changes from liquid to vapour is called (a) evaporation (b) freezing (c) sublimation (d) melting (e) vaporization
3. The constant temperature at which a substance changes state from liquid to solid is called (a) freezing point (b) melting point (c) boiling point (d) heating point
4. Melting point is a temperature at which a solid substance has its bond broken such that it now flows as liquid. (a) freezing point (b) melting point (c) boiling point (d) heating point
5. What effect will reducing the surface pressure of a liquid have on its boiling point? (a) increase its boiling point (b) reduces its boiling point (c) it does not have effect (d) increases its boiling point by 5°C (e) none of the above
6. Which of the following statement is true about evaporation (a) evaporation occurs at all temperature except at boiling points (b) evaporation only occurs at boiling point (c) evaporation occurs at all temperature (d) all of the above (e) none of the above
7. Impurities change the boiling points of liquid. So, salt added to water will (a) increase its boiling point (b) reduces its boiling point (c) it does not have effect (d) reduces its boiling point by 5°C (e) none of the above
8. Which of the following is the correct SI unit of specific latent heat? (a) J/kg (b) J/kg·K (c) J/kg·K (d) J/kg (e) J/kg
9. The heat required to change 1kg of ice at 0°C to water at the same temperature is called (a) specific latent heat of vaporization (b) specific heat capacity (c) specific latent heat of fusion (d) heat capacity (e) specific heat capacity
10. The equation $PV^{\gamma}T^{\gamma} = \text{constant}$ is Boyle's law if (a) $x=0, y=0, z=1$ (b) $x=1, y=0, z=0$ (c) $x=1, y=1, z=0$ (d) $x=1, y=1, z=1$ (e) $x=1, y=1, z=-1$
11. A fixed mass of gas of volume 600cm³ at a temperature of 27°C is cooled at constant pressure to a temperature of 0°C. What is the change in volume? (a) 54cm³ (b) 273cm³ (c) 300cm³ (d) 546cm³ (e) 600cm³
12. A fixed mass of gas occupies a volume of 20cm³ at a fixed pressure of 700mmHg. Assuming that the temperature remains constant, what will be the volume of the gas at 750mmHg? (a) 2.5cm³ (b) 15.5cm³ (c) 18.7cm³ (d) 21.4cm³ (e) 72.5cm³
13. _____ is an agent that changes or tends to change the state of rest or of uniform motion on a straight line of a body (a) force (b) motion (c) friction (d) movement
14. _____ is the movement of an object from one place to another at a particular time (a) force (b) motion (c) friction (d) movement
15. Force that are in contact with the body they affect are called _____ (a) contact force (b) force field (c) non-contact force (d) gravitational force
16. Force whose sources does not required contact with the body on which they act is called _____ (a) contact force (b) force field (c) non-contact force (d) gravitational force
17. If 1.13×10^4 J of heat energy is required to convert 15kg of steam to water. Calculate the specific latent heat of vaporization of steam (a) 7533J/kg (b) 7500J/kg (c) 7870J/kg (d) 8245J/kg

18. The law that states that the pressure of a fixed mass of gas varies inversely as the volume at constant temperature is called _____ (a) Boyle's law (b) Charles's law (c) pressure law (d) general gas law
19. _____ states that for a fixed mass of gas at constant pressure, the volume is proportional to its absolute temperature (a) Boyle's law (b) Charles's law (c) pressure law (d) general gas law
20. When two bodies collide and then move with different velocities (a) elastic collision (b) inelastic collision (c) momentum (d) displacement
21. Which of the following is not an effect of heat? A. expansion B. contraction C. change of state. D. increase in weight
22. Temperature can be measured in the following units except A. degree Celsius B. kelvin C. degree Fahrenheit D. centigrade-meter
23. When heat is extracted from water it changes from liquid to gas. What is this process called _____ (a) vaporization (b) evaporation (c) condensation
24. The process of by which a gas is converted to a solid is called _____ (a) sublimation (b) deposition (c) vaporize
25. Fahrenheit is a unit for _____ (a) heat (b) gas (c) temperature (a) water
26. _____ of a substance is defined as the amount of heat required to raise the temperature of 1 gram of a substance through 1°C. (a) special heat capacity (b) specific heat capacity (c) molar heat capacity
27. _____ of a substance is defined as the amount of heat required to raise the temperature of 1 mole of a substance through 1° C (a) special heat capacity (b) specific heat capacity (c) molar heat capacity
28. There are _____ types of heat capacities (a) 2 (b) 4 (c) 3
29. _____ The heat supplied to a system to raise its temperature through 1° C keeping the volume of the system constant is called (a) heat capacity at constant pressure (b) heat capacity at constant volume (c) thermodynamics
30. The heat supplied to a system to raise its temperature through 1° C keeping the external pressure constant is called _____ (a) heat capacity at constant pressure (b) heat capacity at constant volume (c) thermodynamics
31. An electric heater rated 12V is used to heat 450g of water when a current of 5A was passed through it. What is the final temperature after 30minutes? [Specific capacity of water = 4200 J/kg·K] (a) 57k (b) 20k (c) 57k (d) 80k (e) 40k
32. If the pressure of the vapor on top of an enclosed liquid is equal to the atmospheric pressure, what will be the temperature of the liquid enclosed? (a) room temperature (b) boiling point (c) freezing point (d) standard temperature
33. 200g of water at 90°C is mixed with 100g of water at 30°C. What is the final temperature? (a) 50°C (b) 60°C (c) 70°C (d) 80°C
34. Which of the following is not used to determine the heat content of a body? (a) mass of the body (b) volume of the body (c) specific heat capacity of the body (d) temperature of the body (e) all of the above
35. A _____ is an instrument used in measuring the quantity of heat (a) Barometer (b) Aerometer (c) anemometer (d) calorimeter
36. _____ is defined as the process of reducing heat hot to the surrounding (a) lagging (b) calorimeter (c) regnant (d) capacity
37. It is the spectrum of sunlight in nature. It is formed due to the dispersion of sunlight

Mock Exams Ss2 Syllable

Richard Bailey



Mock Exams Ss2 Syllable:

This book delves into Mock Exams Ss2 Syllable. Mock Exams Ss2 Syllable is a crucial topic that needs to be grasped by everyone, ranging from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Mock Exams Ss2 Syllable, encompassing both the fundamentals and more intricate discussions.

1. This book is structured into several chapters, namely:

- Chapter 1: Introduction to Mock Exams Ss2 Syllable
- Chapter 2: Essential Elements of Mock Exams Ss2 Syllable
- Chapter 3: Mock Exams Ss2 Syllable in Everyday Life
- Chapter 4: Mock Exams Ss2 Syllable in Specific Contexts
- Chapter 5: Conclusion

2. In chapter 1, this book will provide an overview of Mock Exams Ss2 Syllable. The first chapter will explore what Mock Exams Ss2 Syllable is, why Mock Exams Ss2 Syllable is vital, and how to effectively learn about Mock Exams Ss2 Syllable.
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4. In chapter 3, the author will examine the practical applications of Mock Exams Ss2 Syllable in daily life. The third chapter will showcase real-world examples of how Mock Exams Ss2 Syllable can be effectively utilized in everyday scenarios.
5. In chapter 4, the author will scrutinize the relevance of Mock Exams Ss2 Syllable in specific contexts. The fourth chapter will explore how Mock Exams Ss2 Syllable is applied in specialized fields, such as education, business, and technology.
6. In chapter 5, the author will draw a conclusion about Mock Exams Ss2 Syllable. The final chapter will summarize the key points that have been discussed throughout the book.

The book is crafted in an easy-to-understand language and is complemented by engaging illustrations. This book is highly recommended for anyone seeking to gain a comprehensive understanding of Mock Exams Ss2 Syllable.

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