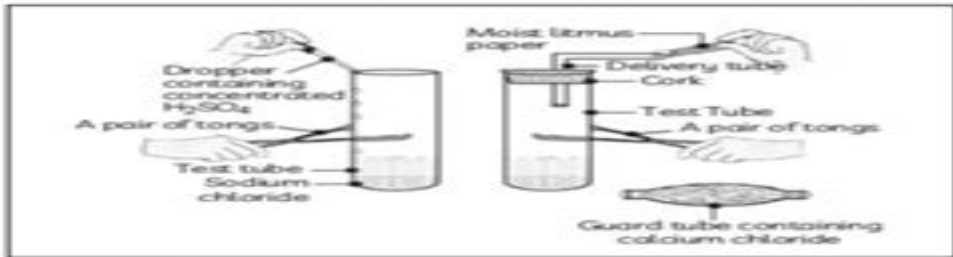
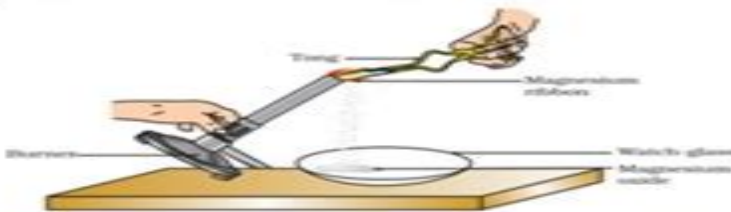


SECTION - A

Select and write one most appropriate option out of the four options given for each of the questions 1 – 20

Q. No	Questions	Marks																									
1	The change in colour of the moist litmus paper in the given set up is due to  i. presence of acid ii. presence of base iii. presence of H⁺(aq) in the solution iv. presence of Litmus which acts as an indicator (a) i and ii (b) Only ii (c) Only iii (d) Only iv.	1																									
2	In the redox reaction $MnO_2 + 4HCl \rightarrow MnCl_2 + 2H_2O + Cl_2$ (a) MnO_2 is reduced to $MnCl_2$ & HCl is oxidized to H_2O (b) MnO_2 is reduced to $MnCl_2$ & HCl is oxidized to Cl_2 (c) MnO_2 is oxidized to $MnCl_2$ & HCl is reduced to Cl_2 (d) MnO_2 is oxidized to $MnCl_2$ & HCl is reduced to H_2O	1																									
3	 Which of the following is the correct observation of the reaction shown in the above set up? (a) Brown powder of Magnesium oxide is formed. (b) Colourless gas which turns lime water milky is evolved. (c) Magnesium ribbon burns with brilliant white light. (d) Reddish brown gas with a smell of burning Sulphur has evolved.	1																									
4	With the reference to four gases CO_2, CO, Cl_2 and O_2, which one of the options in the table is correct? <table><tr><th>Option</th><th>Acidic oxide</th><th>Used in treatment of water</th><th>Product of respiration</th><th>Product of incomplete combustion</th></tr><tr><td>(a)</td><td>CO</td><td>Cl_2</td><td>O_2</td><td>CO</td></tr><tr><td>(b)</td><td>CO_2</td><td>Cl_2</td><td>CO_2</td><td>CO</td></tr><tr><td>(c)</td><td>CO_2</td><td>O_2</td><td>O_2</td><td>CO_2</td></tr><tr><td>(d)</td><td>CO</td><td>O_2</td><td>CO_2</td><td>CO_2</td></tr></table>	Option	Acidic oxide	Used in treatment of water	Product of respiration	Product of incomplete combustion	(a)	CO	Cl_2	O_2	CO	(b)	CO_2	Cl_2	CO_2	CO	(c)	CO_2	O_2	O_2	CO_2	(d)	CO	O_2	CO_2	CO_2	1
Option	Acidic oxide	Used in treatment of water	Product of respiration	Product of incomplete combustion																							
(a)	CO	Cl_2	O_2	CO																							
(b)	CO_2	Cl_2	CO_2	CO																							
(c)	CO_2	O_2	O_2	CO_2																							
(d)	CO	O_2	CO_2	CO_2																							

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