

First/Second Semester B.E. Degree Examination, Dec.2013/Jan.2014
Engineering Chemistry

Time: 3 hrs.

Max. Marks:100

- Notes: 1. Answer any FIVE full questions, choosing at least two from each part.
 2. Answer all objective type questions only on OMR sheet page 3 of the answer booklet.
 3. Answer to objective type questions on sheets other than OMR will not be valued.

PART – A

1. a. Choose the correct answers for the following : (04 Marks)
 - i) A metal rod is dipped in a solution of its ions. Its electrode potential is independent of:

A) temperature of solution	B) concentration of solution
C) area of the metal exposed	D) nature of metal
 - ii) The emf of a cell consisting of a SHE and a metal is found to be 0.74V. The SHE is the positive electrode in the combination. Then the potential of the metal electrode is:

A) +0.74V	B) -1.74V	C) +0.37V	D) -0.74V
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 - iii) Electrode potential of a metal in a dilute solution is:

A) same as that in a concentrated solution
B) lower than that in concentrated solution
C) higher than that in a concentrated solution
D) none of these.
 - iv) The potential of the two metal electrodes used in a cell are 0.55V and 0.85V. For the cell formed by combining them is:

A) 1.20V	B) 0.30V	C) -0.50V	D) -1.20V.
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 - b. What are concentration cell? Explain the working of a concentration cell by taking suitable example. (05 Marks)
 - c. Explain a method for the determination of single electrode potential. (05 Marks)
 - d. An electrochemical cell consists of magnesium electrode in 0.042m $\text{mg}(\text{NO}_3)_2$ solution and silver electrode in 0.35m AgNO_3 solution. The SEP of Mg and Ag are -2.363V and +0.80V respectively. Represent the cell, write the cell reaction and calculate the emf of the cell. (06 Marks)
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2. a. Choose the correct answers for the following : (04 Marks)
 - i) In lead acid battery the product formed on both anode and cathode is:

A) PbO_2	B) PbO	C) PbSO_4	D) Pb
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 - ii) A battery in which a key component is separated from the battery prior to its activation is called:

A) Primary battery	B) Secondary battery	C) Tertiary battery	D) Reserve battery
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 - iii) In which of these batteries aqueous KOH is used as an electrolyte?

A) Ni-cd	B) Ni-MH	C) Zn-air	D) All of these
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 - iv) In which of the following battery the cell reaction is not reversible?

A) Pb- PbO_2	B) Li-MnO ₂	C) Ni-MH	D) Ni-Cd
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 - b. Explain the construction and working of Zn-MnO₂ battery. (05 Marks)
 - c. What are fuel cells? How is differ from battery? Explain the construction and working of $\text{CH}_3\text{OH}-\text{O}_2$ fuel cell. (07 Marks)
 - d. Write the discharging and charging reactions in the following batteries:

i) Ni-Cd battery;	ii) Ni-MH battery.
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(04 Marks)



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