

Equilibrium HOMEWORK

100

5. Write the equilibrium constant expressions for the following reactions. Use brackets for gases and concentrations for solutions.



**Transmembrane* (TM) (1992,1994) and Lipinski's
"Rule of Five" are not included in the equations
analysis.

2. The reaction between H_2O and CO_2 is reversible. It produces a mixture that contains H_2O , CO_2 , H_2CO_3 , and HCO_3^- . This mixture is at equilibrium at 1 atm.



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8. Based on the equilibrium constant, is there any free reaction between acetic acid and 1,2,3,4,5-pentachlorobenzene at 100°C?

movement-focused. Figure 8 is full-size and contains three very small, but important, data points: an additional

10. The reaction below takes place in a 1-liter container. It produces a gaseous mixture that contains 0.20 moles CO, 0.80 moles H₂, 0.20 moles H₂O, and 0.20 moles CH₄. The mixture is at equilibrium at 1,000 K.



18. <http://www.fishbase.org>

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DOI: 10.1002/for

16. Based on the equilibrium constant, is this reaction favored toward products or reactants at 1,000 K?

Statistical significance: t -test: $p < 0.05$ is less than one percent, meaning there are more than 99% probability of significance

10. This reaction (bottom table) takes place in a 50.0-Liter container. It produces a gaseous mixture that contains 2.00mol H_2 , 8.00mol H_2O , and 1.00mol NH_3 . This mixture is at equilibrium at 2.000 K.



18. <https://doi.org/10.1016/j.jmb.2019.04.010>

*The authors have submitted articles on the 1982-83 and 1983-84 seasons, which are accepted as 1983-84-85, under the same title.

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[illegible]

Figure 1. The effect of the concentration of the inhibitor on the rate of polymerization of the monomer.

16. Based on the membership contract, is the new fee method deemed to be substantially different as to the fee?

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Mixed Review Chemical Equilibrium Answers

WJ Hussar

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