

CHEMISTRY



Periodic Table of Elements

1 H 1.008																	18 Ar 39.948		
2 He 4.0026																	36 Kr 83.80		
3 Li 6.941	4 Be 9.0122															54 Xe 131.29	86 Rn 222		
5 Na 22.990	6 Mg 24.305															72 Hf 178.49	84 Po 209		
9 F 18.998	10 Ne 20.180															80 Hg 200.59	118 Og 294		
11 K 39.098	12 Ca 40.078	Sc 44.956	Ti 47.88	V 50.942	Cr 52.00	Mn 54.938	Fe 55.845	Co 58.933	Ni 58.69	Cu 63.546	Zn 65.38	Ga 69.723	Ge 72.64	As 74.922	Se 78.96	Br 79.904	82 Pb 207.2		
13 Al 26.982	14 Si 28.086	21 Sc 44.956	22 Ti 47.88	23 V 50.942	24 Cr 52.00	25 Mn 54.938	26 Fe 55.845	27 Co 58.933	28 Ni 58.69	29 Cu 63.546	30 Zn 65.38	31 Ga 69.723	32 Ge 72.64	33 As 74.922	34 Se 78.96	35 Br 79.904	81 Tl 204.38		
15 P 30.974	16 S 32.06	26 Fe 55.845	27 Co 58.933	28 Ni 58.69	29 Cu 63.546	30 Zn 65.38	31 Ga 69.723	32 Ge 72.64	33 As 74.922	34 Se 78.96	35 Br 79.904	36 Kr 83.80	52 Te 127.6	53 I 126.90	54 Xe 131.29	85 At 210	117 Ts 294		
17 Cl 35.453	18 Ar 39.948	39 Y 88.906	40 Zr 91.224	41 Nb 92.906	42 Mo 95.94	43 Tc 98.906	44 Ru 101.07	45 Rh 102.91	46 Pd 106.42	47 Ag 107.87	48 Cd 112.41	49 In 114.82	50 Sn 118.71	51 Sb 121.76	52 Te 127.6	55 Cs 132.91	119 Nh 294		
19 K 39.098	20 Ca 40.078	37 Rb 85.468	38 Sr 87.62	45 Rh 102.91	46 Pd 106.42	47 Ag 107.87	48 Cd 112.41	49 In 114.82	50 Sn 118.71	51 Sb 121.76	52 Te 127.6	53 I 126.90	54 Xe 131.29	56 Ba 137.33	57 La 138.91	87 Fr 223	120 Ds 294		
25 Mn 54.938	26 Fe 55.845	35 Br 79.904	36 Kr 83.80	43 Tc 98.906	44 Ru 101.07	45 Rh 102.91	46 Pd 106.42	47 Ag 107.87	48 Cd 112.41	49 In 114.82	50 Sn 118.71	51 Sb 121.76	52 Te 127.6	53 I 126.90	54 Xe 131.29	86 Rn 222	118 Og 294		
31 Ga 69.723	32 Ge 72.64	47 Ag 107.87	48 Cd 112.41	49 In 114.82	50 Sn 118.71	51 Sb 121.76	52 Te 127.6	53 I 126.90	54 Xe 131.29	55 Cs 132.91	56 Ba 137.33	57 La 138.91	58 Ce 140.12	59 Pr 140.91	60 Nd 144.24	88 Ra 226	116 Lv 294		
37 Rb 85.468	38 Sr 87.62	55 Cs 132.91	56 Ba 137.33	57 La 138.91	58 Ce 140.12	59 Pr 140.91	60 Nd 144.24	61 Pm 144.91	62 Sm 150.36	63 Eu 151.96	64 Gd 157.25	65 Tb 158.93	66 Dy 162.50	67 Ho 164.93	68 Er 167.26	89 Ac 227	115 Nh 294		
41 Nb 92.906	42 Mo 95.94	61 Pm 144.91	62 Sm 150.36	63 Eu 151.96	64 Gd 157.25	65 Tb 158.93	66 Dy 162.50	67 Ho 164.93	68 Er 167.26	69 Tm 168.93	70 Yb 173.05	71 Lu 174.97	72 Hf 178.49	73 Ta 180.95	74 W 183.84	90 Th 232.04	114 Fl 294		
47 Ag 107.87	48 Cd 112.41	75 Re 186.21	76 Os 190.23	77 Ir 192.22	78 Pt 195.08	79 Au 196.97	80 Hg 200.59	81 Tl 204.38	82 Pb 207.2	83 Bi 208.98	84 Po 209	85 At 210	86 Rn 222	91 Pa 231.04	92 U 238.03	93 Np 237.05	113 Nh 294		
53 I 126.90	54 Xe 131.29	83 Bi 208.98	84 Po 209	85 At 210	86 Rn 222	91 Pa 231.04	92 U 238.03	93 Np 237.05	94 Pu 244.06	95 Am 243.06	96 Cm 247.07	97 Bk 247.07	98 Cf 251.08	99 Es 252.08	100 Fm 257.10	101 Md 258.10	102 No 259.10	103 Lr 262.11	
59 Pr 140.91	60 Nd 144.24	95 Am 243.06	96 Cm 247.07	97 Bk 247.07	98 Cf 251.08	99 Es 252.08	100 Fm 257.10	101 Md 258.10	102 No 259.10	103 Lr 262.11	104 Rf 261.10	105 Db 262.11	106 Sg 266.12	107 Bh 264.10	108 Hs 277.13	109 Mt 268.11	110 Ds 271.10	111 Nh 286.10	112 Nh 289.10

1	ATOMIC NUMBER
H	SYMBOL
1.008	ATOMIC WEIGHT

- Elemental forms:**
 - Monoatomic:** Noble gases: helium, neon, argon, krypton, xenon, and radon; nonreactive
 - Diatomic:** Gases: fluorine, chlorine, hydrogen, oxygen, and nitrogen; *Liquid:* bromine; *Solid:* iodine
 - Molecular solids:** Sulfur
 - Covalent solids:** Carbon (diamond, graphite)
 - Metals:** Most common type of element; high coordination numbers; very dense
- Types of compounds:** Large variety due to ionic, covalent, and metallic bonding options
 - Ionic:** Electron exchange
 - Covalent:** Shared electrons
 - Polar covalent: Unequal sharing
 - Ion complex:** Central ion + solvent
 - Transition metal compounds:**
 - Coordination compounds: Non-bonding ligands
 - Organometallics: Metal-C bond
 - Complex minerals:** Metal oxides and silicates
- Periodicity:** The chemical properties of an element depend on valence electrons
 - Family of elements:** Column in periodic table
 - The addition of d-orbitals produces an expanded octet

Periodic Trends

- Chemistry determined by the behavior of valence electrons
- Atoms or ions get larger as you go down the column; electron shells are added
- Cation and anions (same charge) decrease in size as the atomic number (Z) increases

- Ionization energies:** Noble gases are the most stable; alkali metals are the easiest to ionize; electron screening lessens Z -effect on orbital energy
- Atomic and ionic radii:** Not real properties; derive from structural data; divide bond into radii using a consistent scheme
 - Ionic radius:** Depends on charge, counter ion, and coordination (crystal type)
 - Atomic radius:** Based on diatomic gas data or metallic solid structural data

Development of the Periodic Table

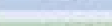
- Missing elements were predicted using periodicity
- Column 1:** Alkali metals
- Column 2:** Alkaline earth metals
- First transition series:** Titanium-Copper
- Column 17:** Halogens: fluorine, chlorine, bromine, iodine, and astatine
- Column 18:** Rare gases: helium, neon, argon, krypton, xenon, and radon
- Actinides:** $Z = 90$ or above; radioactive
- Lanthanides:** f-subshell elements; starts with $Z = 58$



Practical Application

- The colors of fireworks depend on the elements in the explosive

d-subshell
5 orbitals



Atomic Structure

- Atomic number (Z):** The number of protons in the nucleus; for a neutral atom, Z = number of electrons; for an anion of charge $-q$, $Z = q$ electrons; for a cation of charge $+q$, $Z = q$ electrons
- Mass number (A):** $A = Z + N$, where N = the number of neutrons in the nucleus
- Isotopes:** Atoms with the same Z , different A ; an element can have one or more isotopes; any sample contains a number of isotopes; in practical work, use the average mass number or atomic weight (weighted average of natural isotopes for a given element)



Atomic Quantum Numbers & Orbitals

- n : principle**
 - n : magnetic (orbital direction)
 - n : electron spin
- For each n , the possible l values are $0, 1, \dots, n-1$
- For each l , the possible m_l values are $-l, \dots, 0, \dots, +l$
- m_s has two possible values: $+\frac{1}{2}$ and $-\frac{1}{2}$ (spin up, spin down)
- Each value of n denotes a "shell" in the atomic structure; each l denotes a subshell
 - s-orbital: $l = 0$, 1 type
 - d-orbital: $l = 2$, 5 types
 - p-orbital: $l = 1$, 3 types
 - f-orbital: $l = 3$, 7 types

Many-Electron Atoms

Orbitals and quantum numbers derived for the hydrogen atom are used to describe all many-electron atoms and ions; a more rigorous treatment determines exact energy levels and orbital properties; the main result is that subshells have different energies; the filling of levels is guided by the Aufbau principle:

Aufbau Order for Filling Subshells



n	l	m_l	m_s	element
1	0	0	$+\frac{1}{2}$	H
1	0	0	$-\frac{1}{2}$	He
2	0	0	$+\frac{1}{2}$	Li
2	0	0	$-\frac{1}{2}$	Be
2	1	-1	$+\frac{1}{2}$	B
2	1	0	$+\frac{1}{2}$	C
2	1	1	$+\frac{1}{2}$	N
2	1	-1	$-\frac{1}{2}$	O
2	1	0	$-\frac{1}{2}$	F
2	1	1	$-\frac{1}{2}$	Ne

- Electron configuration:** Atomic orbital occupancy for an atom
- Pauli exclusion principle:** Each electron has a unique set of quantum numbers; an orbital may hold up to two electrons: one with spin up, $m_s = +\frac{1}{2}$, and one with spin down, $m_s = -\frac{1}{2}$
- Hund's rule:** Electrons fill p, d, and f subshells in a manner to maximize the number of unpaired electrons (half-filled orbitals); spin is maximized
- Ionization potential (IP):** Energy required to remove electron from an atom or ion
 - First IP:** Removal of first valence electron
- Electronegativity:** Tendency of an atom to attract electrons in a chemical bond; 0-4 scale, 0 for rare gas

Lavc Chemistry Study Guide

Scott C. Dulebohn



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Table of Contents Lavc Chemistry Study Guide

1. Understanding the eBook Lavc Chemistry Study Guide
 - The Rise of Digital Reading Lavc Chemistry Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Lavc Chemistry Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Lavc Chemistry Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Lavc Chemistry Study Guide
 - Personalized Recommendations
 - Lavc Chemistry Study Guide User Reviews and Ratings
 - Lavc Chemistry Study Guide and Bestseller Lists

5. Accessing Lavc Chemistry Study Guide Free and Paid eBooks
 - Lavc Chemistry Study Guide Public Domain eBooks
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 - Lavc Chemistry Study Guide Budget-Friendly Options
6. Navigating Lavc Chemistry Study Guide eBook Formats
 - ePub, PDF, MOBI, and More
 - Lavc Chemistry Study Guide Compatibility with Devices
 - Lavc Chemistry Study Guide Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Lavc Chemistry Study Guide
 - Highlighting and Note-Taking Lavc Chemistry Study Guide
 - Interactive Elements Lavc Chemistry Study Guide
8. Staying Engaged with Lavc Chemistry Study Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Lavc Chemistry Study Guide
9. Balancing eBooks and Physical Books Lavc Chemistry Study Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Lavc Chemistry Study Guide
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Lavc Chemistry Study Guide
 - Setting Reading Goals Lavc Chemistry Study Guide
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Lavc Chemistry Study Guide
 - Fact-Checking eBook Content of Lavc Chemistry Study Guide
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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web solution if you can t see the sugar how can you tell that it is there does sugar dissolve more easily in hot water or cold water gizmo w arm up a solution generally consists of two parts a solute that is dissolved and a solvent that the solute is dissolved into for example sugar is a solute that is dissolved into the solvent water in the

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web 1 in this solution what is the solute potassium nitrate what is the solvent water 2 click add 10 g to mix 10 g of potassium nitrate into the water a did all of the potassium nitrate dissolve b how can you tell yes all of the potassium nitrate disappeared as it was stirred into the water activity a solubility

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