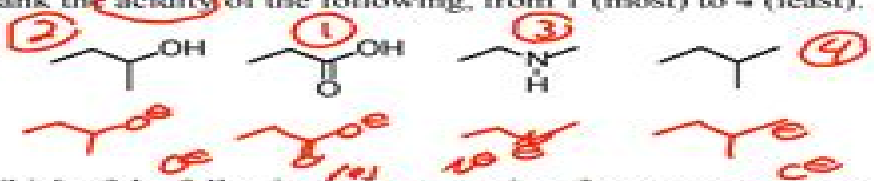


6. Rank the acidity of the following, from 1 (most) to 4 (least). (4pt)

Think Anion Base

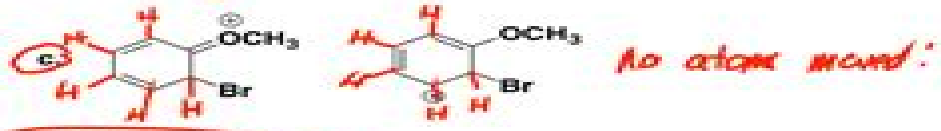


3 Acidity Factors:
 1. charge (not relevant here)
 2. E_{neg}
 3. Resonance

7. Which of the following represent pairs of resonance structures? (4pt)



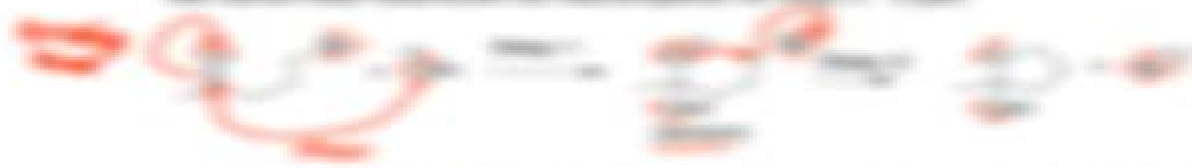
No atoms can move!



d. Both a and c

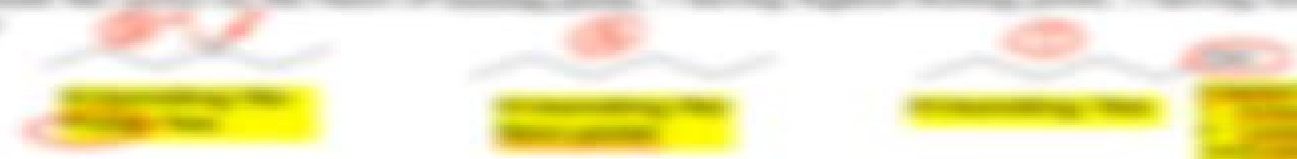
e. a, b, and c are all resonance structures.

8. Draw the Lewis structure for the following molecule, showing all lone pairs and formal charges. (4pt)



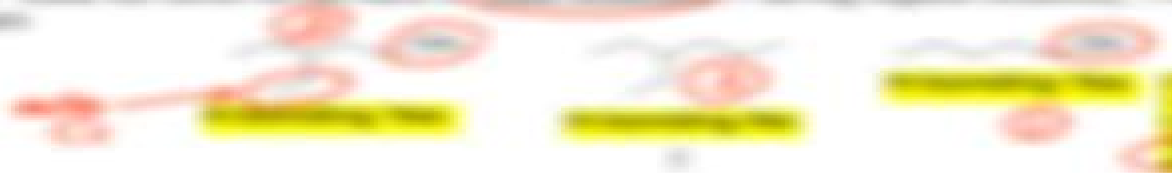
Lewis structure:
 - All atoms must have 8 valence electrons (except H, which has 2).
 - Lone pairs are shown as pairs of dots.
 - Formal charges are calculated as: $FC = V - (L + \frac{B}{2})$

9. Draw the Lewis structure for the following molecule, showing all lone pairs and formal charges. (4pt)



Lewis structure:
 - All atoms must have 8 valence electrons (except H, which has 2).
 - Lone pairs are shown as pairs of dots.
 - Formal charges are calculated as: $FC = V - (L + \frac{B}{2})$

10. Draw the Lewis structure for the following molecule, showing all lone pairs and formal charges. (4pt)



Lewis structure:
 - All atoms must have 8 valence electrons (except H, which has 2).
 - Lone pairs are shown as pairs of dots.
 - Formal charges are calculated as: $FC = V - (L + \frac{B}{2})$

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