

# TOPIC 8 : HEAT & TEMPERATURE

Specific Heat Capacity :-

$$\Delta E = mc\Delta\theta$$

where;  $c$  = specific heat capacity

$m$  = mass

$\Delta\theta$  = change in temperature

Specific Latent Heat :-

$$\Delta E = L\Delta m$$

$L$  = specific latent heat

Kinetic Energy of molecule in a gaseous sample

$$\frac{1}{2}m\langle c^2 \rangle = \frac{3}{2}kT$$

where;  $k$  = Boltzmann constant  
 $1.38 \times 10^{-23} \text{ J K}^{-1}$

$T$  = temperature in Kelvin

## GIAS LAWS

•  $V \propto T$

$$\frac{V_1}{T_1} = \frac{V_2}{T_2}$$

•  $p \propto T$

•  $p \propto \frac{1}{V}$   $p_1V_1 = p_2V_2$

where,  $p$  = pressure

$T$  = temp.

$V$  = vol.

Gras laws ultimate equation

$$pV = NkT$$

where;  $N$  = number of molecules in gas

$k$  = Boltzmann const.

$T$  = temp. in Kelvin

Equation of state for ideal gas

$$pV = nRT$$

$$pV = \frac{1}{3}mN\langle c^2 \rangle$$

where;  $n$  = number of moles of gas

$R$  = universal gas constant =  $8.31 \text{ J kg}^{-1} \text{ mol}^{-1}$

# Jan 2015 Ial Unit 5 Physics Paper

**RS Peters**



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