

The Hydrogen Atom – Student Guide

Background Material

Carefully read the background pages created through Eureka, Slides, and Animations and answer the following questions in detail and understanding.

Question 1: Complete the following table which compares how the Bohr model and the Quantum model compare the Hydrogen atom. In each case, they both describe things in a unique way and in some ways they do not.

Concept Name	Bohr Model	Quantum Model
Electrons	positively charged particles in the atom	positively charged particles in the atom
Orbitals	negatively charged and orbit the nucleus	negatively charged clouds
Location of the electron in the atom without	is in two fixed paths	electron cloud pattern represents the probability
Energy of the electron levels	discrete (separate) the Bohr model	
Energy of the electron levels	$E = -R_H/n^2$	$E = -R_H/n^2$

Question 2: Complete the following table which shows the comparison of two different theories by putting predictions made by each theory side-by-side. The table was completed for you. To see what each is saying, carefully read the responses, consider energy, and the same activity through-out the whole process.

Hydrogen Atom Student Guide Naap Answers

Ensheng Dong



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