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Second Edition

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web correct answer c both a and b q2 the density of 1 m solution of hcl is 1 0585 g ml the molality of the solution is 1 0585 1 0 10 0 0585 correct answer b 1 explanation molality
$$m = \frac{\text{moles of solute}}{\text{kg of solvent}}$$

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web sep 13 2023 the only difference between molality and molarity is in the denominator molality is measured in kilograms of solvent whereas molarity is measured in liters of solution when studying the properties of solutions related to vapour pressure and temperature changes concentrations expressed in molality are used

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web 1 molarity mv mass molar mass x 0 4500 l 0 825 g 141 9579 g mol x 0 0129 m 2 molality 0 825 g 141 9579 g mol 0 00581158 mol 0 00581158 mol 0 4500 kg 0 0129 m 3 mole fraction na 2 hpo 4 0 825 g 141 9579 g mol 0 00581158 mol h 2 o 450 0 g 18 015 g mol 24 97918401 mol

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web apr 7 2014 molarity moles of solute liters of solution 8 4 2 2 a first convert 250 ml to liters 250 1000 0 25 then calculate molarity 5 moles 0 25 liters 20 m 3 c a solution with molarity 2 requires 2 m of n a oh per liter so 4 x 2 8 m 4 a a solution of molarity 1 5 m requires 1 5 mol of na to every litre of solvent

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