



Elements & Macromolecules in Organisms

Most common elements in living things are carbon, hydrogen, nitrogen, and oxygen. These four elements constitute about 95% of your body weight. All compounds can be classified in two broad categories --- organic and inorganic compounds. Organic compounds are made primarily of carbon. Carbon has four outer electrons and can form four bonds. Carbon can also bond to other carbon molecules forming double, triple, or quadruple bonds. Organic compounds also contain hydrogen. Since hydrogen has only one electron, it can form only single bonds.

Each small organic molecule can be a unit of a large organic molecule called a macromolecule. There are four classes of macromolecules (polysaccharides or carbohydrates, triglycerides or lipids, polypeptides or proteins, and nucleic acids such as DNA & RNA). Carbohydrates and lipids are made of only carbon, hydrogen, and oxygen (CHO). Proteins are made of carbon, hydrogen, oxygen, and nitrogen (CHON). Nucleic acids such as DNA and RNA contain carbon, hydrogen, oxygen, nitrogen, and phosphorus (CHON P).

The body also needs trace amounts of other elements such as calcium, potassium, and sulfur for proper functioning of muscles, nerves, etc. *Color* each of the elements according to the color listed next to the element's symbol. *Color code* the squirrel & carrot with the correct proportion of each element's color. *Use the diagrams of glucose* to tell how many carbons, hydrogens, and oxygens are in a single molecule.

Use the drawing of the amino acid on this sheet to determine the number of bonds formed by:

_____ Oxygen

_____ Hydrogen

_____ Nitrogen

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This has been demonstrated by the manipulation of genes that led to the sudden production of additional complete wings and legs in flies and birds. The study of this internal atomic construction of macromolecules is being investigated at the large electron accelerators such as the MAX IV Synchrotron Radiation Laboratory Lund University Sweden. The periodicity of the chemical elements is well known from its iconic Table. Significantly this periodicity can now be seen to extend to the properties of living organisms. Biological properties as different as flight, vision, luminescence and regeneration as well as others show unexpectedly periodic emergence. They resurface without previous announcement in most unrelated plant and animal families and they emerge irrespective of whether the organism is a simple invertebrate or a most complex mammal. Moreover this periodicity does not necessarily start at the cell or DNA levels but appears initially in crystals and minerals where it is shown to be a pure atomic and electronic process e.g. in luminescence and regeneration. The assembled molecular evidence led to the construction of Periodic Tables of living organisms placing them in a position comparable to the periodicity of the chemical elements. Surprisingly there are striking resemblances between the periodicities of the chemical elements and those of living organisms. In addition the two types of Tables increase our insight into the events directing atomic evolution since the periodic law established in chemical elements turns out to be applicable to the periodicity of living organisms. The new Periodic Tables introduce a predictive capacity in biological evolution that before was hardly contemplated. Eric Scerri from the Department of Chemistry and Biochemistry California University Los Angeles who is the Author of the book *The Periodic Table Its Story and its Significance* Oxford University Press stated in an e mail that Professor Lima de Faria's book is wonderful and a pioneering work.

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