

5

ELECTRONS IN ATOMS

A. Matching

Match each term in Column A with the correct description in Column B. Write the letter of the correct term on the line.

Column A

- _____ 1. the ways in which electrons are arranged around the nuclei of atoms
- _____ 2. the ejection of electrons by metals when light shines on them
- _____ 3. the region around the nucleus of an atom where an electron is likely to be moving
- _____ 4. An atomic orbital may describe at most two electrons.
- _____ 5. the regions within which electrons have the highest probability of being found
- _____ 6. When electrons occupy orbitals of equal energy, one electron enters each orbital until all the orbitals contain one electron.
- _____ 7. the amount of energy required to move an electron from its present energy level to the next highest one
- _____ 8. light quanta
- _____ 9. Electrons enter orbitals of lowest energy first.
- _____ 10. the modern description of the electrons in atoms

Column B

- a. Aufbau principle
- b. quantum
- c. photoelectric effect
- d. atomic orbitals
- e. Pauli exclusion principle
- f. energy level
- g. Hund's rule
- h. electron configurations
- i. quantum mechanical model
- j. photons

B. Multiple Choice

Choose the best answer and write its letter on the line.

- _____ 11. Bohr's contribution to the development of atomic structure:
 - a. was referred to as the "planar packing model."
 - b. was the discovery that electrons surround a dense nucleus.
 - c. was proposed that electrons travel in circular orbits around the nucleus.
 - d. is the quantum mechanical model.
- _____ 12. What is the total number of orbitals in the third principal energy level?
 - a. 1
 - b. 4
 - c. 9
 - d. 16

Electrons In Atoms Chapter Test B

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Table of Contents Electrons In Atoms Chapter Test B

1. Understanding the eBook Electrons In Atoms Chapter Test B
 - The Rise of Digital Reading Electrons In Atoms Chapter Test B
 - Advantages of eBooks Over Traditional Books
2. Identifying Electrons In Atoms Chapter Test B
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Electrons In Atoms Chapter Test B
 - User-Friendly Interface
4. Exploring eBook Recommendations from Electrons In Atoms Chapter Test B
 - Personalized Recommendations
 - Electrons In Atoms Chapter Test B User Reviews and Ratings

- Electrons In Atoms Chapter Test B and Bestseller Lists
- 5. Accessing Electrons In Atoms Chapter Test B Free and Paid eBooks
 - Electrons In Atoms Chapter Test B Public Domain eBooks
 - Electrons In Atoms Chapter Test B eBook Subscription Services
 - Electrons In Atoms Chapter Test B Budget-Friendly Options
- 6. Navigating Electrons In Atoms Chapter Test B eBook Formats
 - ePub, PDF, MOBI, and More
 - Electrons In Atoms Chapter Test B Compatibility with Devices
 - Electrons In Atoms Chapter Test B Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Electrons In Atoms Chapter Test B
 - Highlighting and Note-Taking Electrons In Atoms Chapter Test B
 - Interactive Elements Electrons In Atoms Chapter Test B
- 8. Staying Engaged with Electrons In Atoms Chapter Test B
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Electrons In Atoms Chapter Test B
- 9. Balancing eBooks and Physical Books Electrons In Atoms Chapter Test B
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Electrons In Atoms Chapter Test B
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Electrons In Atoms Chapter Test B
 - Setting Reading Goals Electrons In Atoms Chapter Test B
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Electrons In Atoms Chapter Test B
 - Fact-Checking eBook Content of Electrons In Atoms Chapter Test B
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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