

## CHAPTER 4 REVIEW

**Arrangement of Electrons in Atoms****SECTION 2****SHORT ANSWER** Answer the following questions in the space provided.

1. \_\_\_\_\_ How many quantum numbers are used to describe the properties of electrons in atomic orbitals?  
(a) 1 (c) 3  
(b) 2 (d) 4
2. \_\_\_\_\_ A spherical electron cloud surrounding an atomic nucleus would best represent  
(a) an  $s$ -orbital. (c) a combination of two different  $p$ -orbitals.  
(b) a  $p$ -orbital. (d) a combination of an  $s$  and a  $p$ -orbital.
3. \_\_\_\_\_ How many electrons can an energy level of  $n = 4$  hold?  
(a) 32 (c) 8  
(b) 24 (d) 6
4. \_\_\_\_\_ How many electrons can an energy level of  $n = 2$  hold?  
(a) 32 (c) 8  
(b) 24 (d) 6
5. \_\_\_\_\_ Compared with an electron for which  $n = 2$ , an electron for which  $n = 4$  has more  
(a) spin. (c) energy.  
(b) particle nature. (d) wave nature.
6. \_\_\_\_\_ According to Bohr, which is the point in the figure below where electrons cannot reside?  
(a) point A (c) point C  
(b) point B (d) point D



7. \_\_\_\_\_ According to the quantum theory, point D in the above figure represents:  
(a) the fixed position of an electron.  
(b) the furthest position from the nucleus that an electron can achieve.  
(c) a position where an electron probably exists.  
(d) a position where an electron cannot exist.

# Electrons In Atoms Test A Key

**L Darling-Hammond**



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