

## 5

## ELECTRONS IN ATOMS

## A. Matching

Match each term in Column A with the correct description in Column B. Write the letter of the correct term on the line.

## Column A

- \_\_\_\_\_ 1. the ways in which electrons are arranged around the nuclei of atoms
- \_\_\_\_\_ 2. the ejection of electrons by metals when light shines on them
- \_\_\_\_\_ 3. the region around the nucleus of an atom where an electron is likely to be moving
- \_\_\_\_\_ 4. An atomic orbital may describe at most two electrons.
- \_\_\_\_\_ 5. the regions within which electrons have the highest probability of being found
- \_\_\_\_\_ 6. When electrons occupy orbitals of equal energy, one electron enters each orbital until all the orbitals contain one electron.
- \_\_\_\_\_ 7. the amount of energy required to move an electron from its present energy level to the next highest one
- \_\_\_\_\_ 8. light quanta
- \_\_\_\_\_ 9. Electrons enter orbitals of lowest energy first.
- \_\_\_\_\_ 10. the modern description of the electrons in atoms

## Column B

- a. Aufbau principle
- b. quantum
- c. photoelectric effect
- d. atomic orbitals
- e. Pauli exclusion principle
- f. energy level
- g. Hund's rule
- h. electron configurations
- i. quantum mechanical model
- j. photons

## B. Multiple Choice

Choose the best answer and write its letter on the line.

- \_\_\_\_\_ 11. Bohr's contribution to the development of atomic structure:
  - a. was referred to as the "planet parking model."
  - b. was the discovery that electrons surround a dense nucleus.
  - c. was proposed that electrons travel in circular orbits around the nucleus.
  - d. is the quantum mechanical model.
- \_\_\_\_\_ 12. What is the total number of orbitals in the third principal energy level?
  - a. 1
  - b. 4
  - c. 9
  - d. 16

# Electrons In Atoms Chapter 5 Test B Answers

**Jay K. Hackett**



## Electrons In Atoms Chapter 5 Test B Answers:

*Science in Your World: Teacher edition* Jay K. Hackett, 1991      *Prentice Hall Physical Science Concepts in Action Program Planner National Chemistry Physics Earth Science*, 2003-11 Prentice Hall Physical Science Concepts in Action helps students make the important connection between the science they read and what they experience every day Relevant content lively explorations and a wealth of hands on activities take students understanding of science beyond the page and into the world around them Now includes even more technology tools and activities to support differentiated instruction      *Chemical Structure and Reactivity* James Keeler, Peter Wothers, 2013-11 Why do certain substances react together in the way that they do What determines the shape of molecules And how can we predict whether a particular reaction will happen at all Such questions lie at the heart of chemistry the science of understanding the composition of substances their reactions and properties Though introductory chemistry is often broken into three sections inorganic organic and physical the only way for students to fully understand the subject is to see it as a single unified whole Chemical Structure and Reactivity rises to the challenge of depicting the reality of chemistry Offering a fresh approach to the subject by depicting it as a seamless discipline the text shows how organic inorganic and physical concepts can be blended together in order to achieve the common goal of understanding chemical systems With a lively and engaging writing style enhanced by vivid illustrations only Chemical Structure and Reactivity makes teaching chemistry with an integrated approach possible Special Features The only introductory text to take a truly integrated approach in explaining the fundamentals of chemistry Fosters an orbital based understanding of reactions with clear curly arrow mechanistic detail throughout A two part structure allows flexibility of use Part I lays down the core of the subject while Part II describes a series of relatively standalone topics which can be selected to fit a particular course Numerous concepts are illustrated with fully cross referenced custom developed online modules enabling students to develop an understanding through active learning Self test exercises embedded in the text with solutions at the end of each chapter and extensive question sets encourage hands on learning to help students master the subject and gain confidence The Online Resource Centre features a range of additional resources for both students and registered adopters of the book New to this Edition A new chapter on symmetry has been added to Part I Discussions of organometallic chemistry spectroscopy and molecular geometry have been expanded Cross references from Part I to Part II have been increased to make the links between core concepts and more advanced topics clearer More self test questions and exercises have been provided      *Chemical Interactions*, 2005      *Discover Science: Teacher's annotated edition*, 1991 Science content helps develop the skills needed to understand how science works learn new concepts solve problems and make decisions in today's technological society      **Ebook: Biology** BROOKER, 2014-09-16 Ebook Biology      *Disha 30 New Syllabus Chapter-wise, Topic-wise & Skill-wise CBSE Class 12 Chemistry Previous Year Solved Papers (2013 - 2025) & Sample Papers (2022 - 2025) 3rd Edition* | PYQs for 2026 Exam Disha Experts, 2025-09-02 The 3rd updated edition of the

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