

**5****ELECTRONS IN ATOMS****Practice Problems***In your notebook, solve the following problems.***SECTION 5.1 MODELS OF THE ATOM**

- How many sublevels are in the following principal energy levels?
  - $n = 1$
  - $n = 2$
  - $n = 3$
  - $n = 4$
  - $n = 5$
  - $n = 6$
- How many orbitals are in the following sublevels?
  - 1s sublevel
  - 5s sublevel
  - 4d sublevel
  - 4f sublevel
  - 7s sublevel
  - 3p sublevel
  - fifth principal energy level
  - 6d sublevel
- What are the types of sublevels and number of orbitals in the following energy levels?
  - $n = 1$
  - $n = 2$
  - $n = 3$
  - $n = 4$
  - $n = 5$

**SECTION 5.2 ELECTRON ARRANGEMENT IN ATOMS**

- Write a complete electron configuration of each atom.
  - hydrogen
  - vanadium
  - magnesium
  - barium
  - bromine
  - sulfur
  - krypton
  - arsenic
  - radon

**SECTION 5.3 PHYSICS AND THE QUANTUM MECHANICAL MODEL**

- What is the wavelength of the radiation whose frequency is  $5.00 \times 10^{15} \text{ s}^{-1}$ ? In what region of the electromagnetic spectrum is this radiation?
- An inexpensive laser that is available to the public emits light that has a wavelength of 670 nm. What are the color and frequency of the radiation?
- What is the energy of a photon whose frequency is  $2.22 \times 10^{14} \text{ s}^{-1}$ ?
- What is the frequency of a photon whose energy is  $6.00 \times 10^{-19} \text{ J}$ ?
- Arrange the following types of electromagnetic radiation in order of increasing frequency.
  - infrared
  - cosmic rays
  - visible light
  - radio waves
  - microwaves
  - ultraviolet
- Suppose that your favorite AM radio station broadcasts at a frequency of 1600 kHz. What is the wavelength in meters of the radiation from the station?

# Electrons In Atoms Answer Key

**K Morrison**



## **Electrons In Atoms Answer Key:**

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