B

ELECTRONS IN ATOMS

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SECTION 13.1 MODELS OF THE ATOM (pages 361-366)

This section summarizes the development of atomic theory. It also explains the significance of quantized energies of electrons as they relate to the quantum mechanical model of the atom.

➤ The Evolution of Atomic Models (pages 361-363)

1. What are the chemical properties of atoms, ions, and molecules related to?

The properties are related to the arrangement of electrons.

Complete the table about atomic models and the scientists who developed them.

Scientist	Model of Atom
Dalton	The atom is a solid indivisible mass.
Thomson	The atom is a ball of positive charge with electrons stuck into the ball.
Rutherford	Most of an atom's mass is concentrated in the small, positively charged nucleus. The electrons surround the nucleus and the rest of the atom is empty space.
Bohr	Electrons are arranged in concentric circular paths around the nucleus.

- The energy level of an electron is the region around the nucleus where the electron is likely to be moving
- Is the following sentence true or false? The electrons in an atom can exist between energy levels.
- Circle the letter of the term that completes the sentence correctly. A quantum of energy is the amount of energy required to
 - a. move an electron from its present energy level to the next lower one
 - b. maintain an electron in its present energy level
 - c.) move an electron from its present energy level to the next higher one
- 6. In general, the higher the electron is on the energy ladder, the

farther it is from the nucleus.

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L Reisser

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Table of Contents Electrons And Atoms Chapter 5 Workbook Answers

- 1. Understanding the eBook Electrons And Atoms Chapter 5 Workbook Answers
 - The Rise of Digital Reading Electrons And Atoms Chapter 5 Workbook Answers
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Electrons And Atoms Chapter 5 Workbook Answers
 - Exploring Different Genres
 - $\circ\,$ Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Electrons And Atoms Chapter 5 Workbook Answers
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Electrons And Atoms Chapter 5 Workbook Answers
 - Personalized Recommendations
 - Electrons And Atoms Chapter 5 Workbook Answers User Reviews and Ratings
 - Electrons And Atoms Chapter 5 Workbook Answers and Bestseller Lists
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 - Electrons And Atoms Chapter 5 Workbook Answers Budget-Friendly Options
- 6. Navigating Electrons And Atoms Chapter 5 Workbook Answers eBook Formats
 - ∘ ePub, PDF, MOBI, and More
 - Electrons And Atoms Chapter 5 Workbook Answers Compatibility with Devices
 - Electrons And Atoms Chapter 5 Workbook Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Electrons And Atoms Chapter 5 Workbook Answers
 - Highlighting and Note-Taking Electrons And Atoms Chapter 5 Workbook Answers
 - Interactive Elements Electrons And Atoms Chapter 5 Workbook Answers
- 8. Staying Engaged with Electrons And Atoms Chapter 5 Workbook Answers

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Electrons And Atoms Chapter 5 Workbook Answers
- 9. Balancing eBooks and Physical Books Electrons And Atoms Chapter 5 Workbook Answers
 - Benefits of a Digital Library
 - o Creating a Diverse Reading Collection Electrons And Atoms Chapter 5 Workbook Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Electrons And Atoms Chapter 5 Workbook Answers
 - Setting Reading Goals Electrons And Atoms Chapter 5 Workbook Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Electrons And Atoms Chapter 5 Workbook Answers
 - Fact-Checking eBook Content of Electrons And Atoms Chapter 5 Workbook Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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