

Electron Configuration Practice

In the space below, write the standard electron configurations of the following elements:

- 1) nitrogen $1s^2 2s^2 2p^3$
- 2) copper $1s^2 2s^2 2p^6 3s^2 3p^6 4s^1 3d^{10}$
- 3) chlorine $1s^2 2s^2 2p^6 3s^2 3p^5$
- 4) calcium $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2$
- 5) lead $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{14} 5d^{10} 6p^2$

In the space below, write the noble gas notation electron configurations of the following elements:

- 6) barium $[Xe] 6s^2$
- 7) gold $[Xe] 4f^{14} 5d^{10} 6s^1$
- 8) platinum $[Xe] 4f^{14} 5d^9 6s^1$
- 9) xenon $[Kr] 4d^{10} 5s^2 p^6$
- 10) calcium $[Ar] 4s^2$

Determine what elements are denoted by the following electron configurations:

- 11) $1s^2 2s^2 2p^6 3s^2 3p^3$ phosphorous
- 12) $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^1 4d^7$ ruthenium
- 13) $[Kr] 4d^{10} 5s^2 p^5$ iodine
- 14) $[Xe] 4f^{14} 5d^{10} 6s^2$ mercury
- 15) $[Rn] 7s^2$ radium

Determine which of the following electron configurations are not correct:

- 16) $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2$ $3p^5$ should be a 6
- 17) $1s^3 2s^2 2p^6 3s^2 2p^5$ 1s only holds 2 and there aren't two "2p" orbitals
- 18) $[Cs] 5f^{14} 6d^1 7s^2$ Cs is not a noble gas
- 19) $[Kr] 4d^{10} 5s^2 p^4$ Correct
- 20) $[Ne]$ for Neon Can't use a noble gas to shortcut itself

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Elements Of Green Function And Density Functional Theory Ferdi Aryasetiawan, 2025-03-17 If there were no Coulomb interaction among electrons it would be relatively straightforward to solve the many electron Schrödinger equation It is however precisely this interaction that is at the heart of numerous fascinating phenomena in condensed matter physics such as superconductivity Kondo physics magnetism etc Due to the large number of electrons in a material being of the order of Avogadro's number it is at present and perhaps in the foreseeable future not feasible or even desirable to solve the Schrödinger equation to obtain the many electron wavefunction Fortunately a large number of important physical properties can be calculated without explicit knowledge of the wavefunction Two of the most important formalisms for dealing with the many electron problem which avoid a direct use of the many electron wavefunction are the Green function and the density functional theory Within the Kohn Sham scheme the latter is used to calculate ground state properties whereas the former for excitation spectra The book presents the fundamentals of these two theories in detail with essential many body tools such as the occupation number representation and Grassmann algebra developed from scratch Prior knowledge of many body theory is not a prerequisite so that it is readable for final year undergraduates and graduate students in physics and chemistry as well as researchers in the field of electronic structure and many body theory The book includes in the last chapter an exposition of a density functional path for determining the Green function a new formalism recently proposed by the author The book should be a valuable companion for those embarking in the field of many electron physics

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