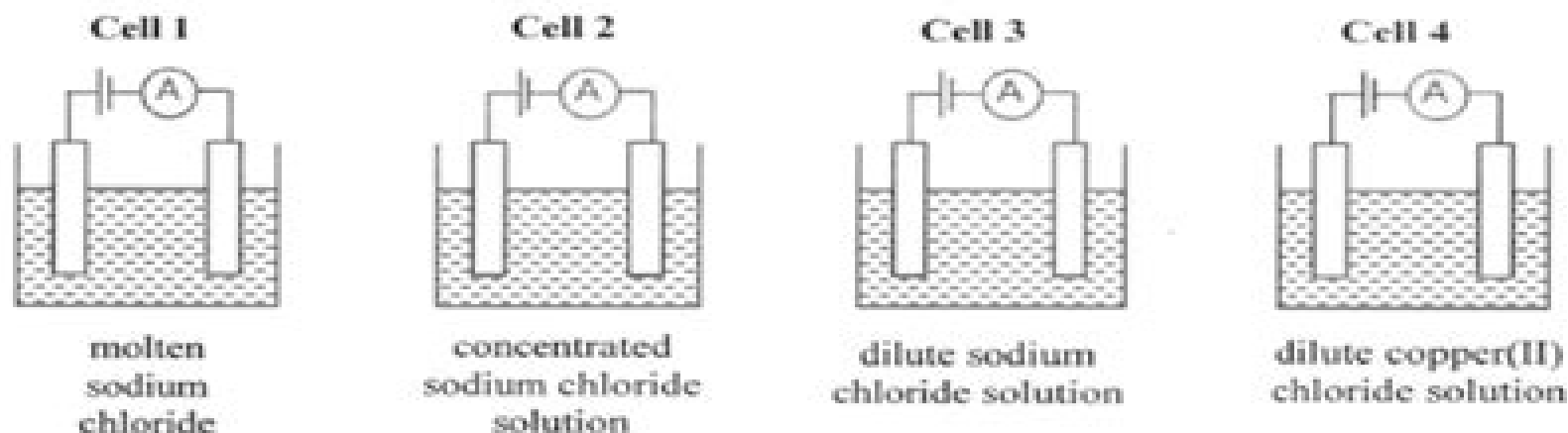


The diagrams show the electrolysis of four different electrolytes, using graphite electrodes.



- (a) Explain why sodium metal can be extracted from molten sodium chloride solution using **Cell 1**. [1]

Sodium ions gains electrons at the cathode and is reduced to form sodium metal.

- (b) The gases produced at the electrodes in **Cell 2** and **Cell 3** are collected and their volumes are measured.

- (i) Name the gas formed at the anode of **Cell 3** and describe a test to identify the gas. [1]

Oxygen gas is formed at the anode of Cell 3.

Insert a glowing splint into a test tube of the gas. The glowing splint will be re-ignited, showing the presence of oxygen gas.

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