

Classifying and Balancing Equations
Multiple Choice

Name _____
Date _____

Per _____

PRACTICE TEST

- _____ 1. A chemical reaction is a process in which
a. products change into reactants. c. substances can change state
b. the law of conservation of mass applies d. all of these
- _____ 2. During a chemical reaction,
a. new elements are produced c. atoms are destroyed
b. atoms are rearranged d. elements are destroyed
- _____ 3. An equation is balanced by
a. changing subscripts c. erasing elements as necessary
b. adding coefficients d. adding elements as necessary
- _____ 4. An atom's ability to undergo chemical reactions is determined by its
a. protons b. innermost electrons c. neutrons d. valence electrons
- _____ 5. What are the reactants in the following chemical equation:
$$\text{Zn} + \text{CuSO}_4 \longrightarrow \text{ZnSO}_4 + \text{Cu}$$

a. zinc and copper c. zinc and copper (II) sulfate
b. zinc sulfate and copper d. only zinc
- _____ 6. What are the products in the above equation?
a. zinc and copper c. zinc and copper (II) sulfate
b. zinc sulfate and copper d. only zinc

Short Answer

For questions 7-12, classify the reaction according to the type it is. Put that answer in the blank. Then add coefficients to balance the reaction when necessary.

- _____ 7. $\text{Zn} + \text{H}_2\text{SO}_4 \longrightarrow \text{ZnSO}_4 + \text{H}_2$
- _____ 8. $\text{H}_2\text{CO}_3 \longrightarrow \text{CO}_2 + \text{H}_2\text{O}$
- _____ 9. $\text{CaCO}_3 + \text{HCl} \longrightarrow \text{CaCl}_2 + \text{H}_2\text{CO}_3$
- _____ 10. $\text{AgNO}_3 + \text{Zn} \longrightarrow \text{Zn}(\text{NO}_3)_2 + \text{Ag}$
- _____ 11. $\text{C}_3\text{H}_8 + \text{O}_2 \longrightarrow \text{CO}_2 + \text{H}_2\text{O}$
- _____ 12. $\text{C}_2\text{H}_5\text{OH} + \text{O}_2 \longrightarrow \text{CO}_2 + \text{H}_2\text{O}$

Write a balanced equation for each of the following reactions:

13. Magnesium chloride is the product of a reaction between magnesium and chlorine.

14. Copper (II) hydroxide and potassium sulfate are produced when potassium hydroxide reacts with copper (II) sulfate.

Unit 6 Practice Test - Key

1. b
2. b
3. b
4. d
5. c
6. B
7. Single replacement 1,1,1,1
8. Decomposition 1,1,1
9. Double replacement 1,2,1,1-
CaCO₃ is ppt (see Ref. Table)
10. Single replacement 2,1,1,2

11. Combustion 1, 5, 3, 4
12. Combustion 1,3,2,3
13. $\text{Mg} + \text{Cl}_2 \rightarrow \text{MgCl}_2$ 1,1,1
14. $\text{KOH} + \text{CuSO}_4 \rightarrow \text{Cu}(\text{OH})_2 + \text{K}_2\text{SO}_4$ 2,1,1,1

Modeling Chemistry Unit 7 Chemical Reactions Key

L Cohen



Modeling Chemistry Unit 7 Chemical Reactions Key:

Decoding **Modeling Chemistry Unit 7 Chemical Reactions Key**: Revealing the Captivating Potential of Verbal Expression

In a period characterized by interconnectedness and an insatiable thirst for knowledge, the captivating potential of verbal expression has emerged as a formidable force. Its ability to evoke sentiments, stimulate introspection, and incite profound transformations is genuinely awe-inspiring. Within the pages of "**Modeling Chemistry Unit 7 Chemical Reactions Key**," a mesmerizing literary creation penned by a celebrated wordsmith, readers attempt an enlightening odyssey, unraveling the intricate significance of language and its enduring affect our lives. In this appraisal, we shall explore the book is central themes, evaluate its distinctive writing style, and gauge its pervasive influence on the hearts and minds of its readership.

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