

Chemistry – Unit 5 Worksheet 2 **Empirical and Molecular Formulation**

Work all your best efforts solving the following problems. Do not forget to show your work and label your answers.

1. Find the empirical formula of a compound containing 33.9 g of bromine and 4.7 g of nitrogen.

$$\text{mole Br} = \frac{33.9 \text{ g Br}}{79.9 \text{ g/mol}} = 0.424 \text{ mole Br}$$

1:1 ratio

$$\text{mole N} = \frac{4.7 \text{ g N}}{14.0 \text{ g/mol}} = 0.336 \text{ mole N}$$



2. What is the empirical formula of a hydrocarbon compound if 1.00 g of this compound contains 0.75 g of carbon and 0.25 g of hydrogen?

$$\text{mole C} = \frac{0.75 \text{ g C}}{12.0 \text{ g/mol}} = 0.0625 \text{ g}$$

$$\text{mole H} = \frac{0.25 \text{ g H}}{1.0 \text{ g/mol}} = 0.25 \text{ mole H}$$

1:4 ratio

$$\text{CH}_4$$



3. A compound was identified and determined to contain 40.0 g of nitrogen, 10.0 g of hydrogen, and 10.0 g of oxygen. What is the empirical formula of this compound?

$$\text{mole N} = \frac{40.0 \text{ g N}}{14.0 \text{ g/mol}} = 2.86 \text{ mole N}$$

$$\text{mole H} = \frac{10.0 \text{ g H}}{1.0 \text{ g/mol}} = 10.0 \text{ mole H}$$

1:3:1 ratio

$$\text{mole O} = \frac{10.0 \text{ g O}}{16.0 \text{ g/mol}} = 0.625 \text{ mole O}$$



4. A compound composed of hydrogen and oxygen is found to contain 10.0 g of hydrogen and 10.0 g of oxygen. What is the ratio of these components if it is given. Find the empirical formula.

$$\text{mole H} = \frac{10.0 \text{ g H}}{1.0 \text{ g/mol}} = 10.0 \text{ mole H}$$

$$\text{mole O} = \frac{10.0 \text{ g O}}{16.0 \text{ g/mol}} = 0.625 \text{ mole O}$$

1:16 ratio hydrogen to oxygen 1:16

simplest form, divide with 0.625

oxygen 16 the oxygen is 16

divide 16 by 16, we have

1 hydrogen 1 oxygen 1:1 H_2O

Modeling Chemistry Unit 5 Review Answers

SA Dillow



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