

Name Tara ShettyDate 9/24/15 Pd 5**Chemistry – Unit 4 Worksheet 3**

Use the following information about the masses of elements in each pair of compounds to help you suggest formulas that account for these ratios.

1. Compounds of carbon and oxygen

Compound A: 57.1 g O / 42.9 g C

Compound B: 72.7 g O and 27.3 C

$$A = \frac{57.1}{42.9} = 1\frac{1}{3}$$

$$B = \frac{72.7}{27.3} = 2\frac{2}{3}$$

a. Determine the value of the ratio $\frac{\text{mass O}}{\text{mass C}}$ in each compound. A $1\frac{1}{3}$ B $2\frac{2}{3}$

b. How does the mass ratio for compound B compare to that in compound A?

compound B is larger than compound A

c. Express these ratios as improper fractions. $\frac{57.1}{42.9} + \frac{72.7}{27.3} \left(\frac{4}{3} + \frac{8}{3} \right)$

d. For each hypothesis, sketch particle diagrams for the compounds of A and B that account for these mass ratios. Write the formula for the compound in each diagram.

Hypothesis 1

Atoms of C and O have the same mass

A $\frac{4}{3}$ B $\frac{8}{3}$ **Hypothesis 2**

Atoms of O are heavier than C atoms by the ratio in compound A.

A



B



Modeling Chemistry Unit 10 2 Answer

Daniel F McAuley



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Table of Contents Modeling Chemistry Unit 10 2 Answer

1. Understanding the eBook Modeling Chemistry Unit 10 2 Answer
 - The Rise of Digital Reading Modeling Chemistry Unit 10 2 Answer
 - Advantages of eBooks Over Traditional Books
2. Identifying Modeling Chemistry Unit 10 2 Answer
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Modeling Chemistry Unit 10 2 Answer
 - User-Friendly Interface
4. Exploring eBook Recommendations from Modeling Chemistry Unit 10 2 Answer
 - Personalized Recommendations

- Modeling Chemistry Unit 10 2 Answer User Reviews and Ratings
- Modeling Chemistry Unit 10 2 Answer and Bestseller Lists
- 5. Accessing Modeling Chemistry Unit 10 2 Answer Free and Paid eBooks
 - Modeling Chemistry Unit 10 2 Answer Public Domain eBooks
 - Modeling Chemistry Unit 10 2 Answer eBook Subscription Services
 - Modeling Chemistry Unit 10 2 Answer Budget-Friendly Options
- 6. Navigating Modeling Chemistry Unit 10 2 Answer eBook Formats
 - ePub, PDF, MOBI, and More
 - Modeling Chemistry Unit 10 2 Answer Compatibility with Devices
 - Modeling Chemistry Unit 10 2 Answer Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Modeling Chemistry Unit 10 2 Answer
 - Highlighting and Note-Taking Modeling Chemistry Unit 10 2 Answer
 - Interactive Elements Modeling Chemistry Unit 10 2 Answer
- 8. Staying Engaged with Modeling Chemistry Unit 10 2 Answer
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Modeling Chemistry Unit 10 2 Answer
- 9. Balancing eBooks and Physical Books Modeling Chemistry Unit 10 2 Answer
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Modeling Chemistry Unit 10 2 Answer
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Modeling Chemistry Unit 10 2 Answer
 - Setting Reading Goals Modeling Chemistry Unit 10 2 Answer
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Modeling Chemistry Unit 10 2 Answer
 - Fact-Checking eBook Content of Modeling Chemistry Unit 10 2 Answer

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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