

Chemistry Unit 8 Worksheet 4

Examples of Every Kind of Problem

This is a separate sheet of paper, so this is a complete solution to each of the problems. You can follow the problems without an answer. We want to make your final solution.

12. Calculate the reaction of sulfur of potassium chlorate, KClO_3 , and potassium chlorate to produce potassium chloride KCl and a certain amount of oxygen gas.

Equation: $2\text{KClO}_3 \rightarrow 2\text{KCl} + 3\text{O}_2$

Mass:	1.0	0	0
Change:	-1.0	+0.5	+1.5
Mass:	0	0.5	1.5

1.0 g of KClO_3 is needed

13. As a single displacement reaction, magnesium metal reacts with hydrochloric acid to produce magnesium chloride and hydrogen gas. How many moles of hydrochloric acid are needed to completely react with 5.0 g of magnesium?

Equation: $\text{Mg} + 2\text{HCl} \rightarrow \text{MgCl}_2 + \text{H}_2$

Mass:	0.200	0	0	0
Change:	-0.200	-0.400	+0.200	+0.200
Mass:	0	0	0.200	0.200

0.400 moles HCl needed

14. Ethane, C_2H_6 , reacts with oxygen gas to produce carbon dioxide gas and water vapor. What mass of oxygen gas is required to react with 1.0 g of ethane?

Equation: $\text{C}_2\text{H}_6 + 3.5\text{O}_2 \rightarrow 2\text{CO}_2 + 3\text{H}_2\text{O}$

Mass:	1.00	0.00	0	0
Change:	-1.00	-2.25	+2.00	+1.50
Mass:	0	2.25	2.00	1.50

2.25 g of O_2 is needed

15. Sodium metal reacts with sulfur dioxide to produce sodium sulfide and sodium sulfate. How many moles of sodium metal are needed to produce 1.0 g of sodium sulfate?

Equation: $4\text{Na} + \text{SO}_2 \rightarrow \text{Na}_2\text{S} + 2\text{Na}_2\text{SO}_4$

Equation: $4\text{Na} + \text{SO}_2 \rightarrow \text{Na}_2\text{S} + 2\text{Na}_2\text{SO}_4$

Mass:	0.000	0.00	0	0
Change:	-0.000	-0.000	+0.000	+0.000
Mass:	0	0.00	0.000	0.000

0.000 g of Na_2S is needed

Modeling Chemistry U8 Ws 3 V2 Key

Anthony S. Fauci



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