For the reverse reaction? <u>(ICC/EaSed</u> The Ea/r)
q) What effect did the catalyst have on the ΔH of the forward reaction? ho effect
The reverse reaction? no effect
12. Name four instances in which catalysts are used in industry or everyday life and tell which catalysts are used.  V20s Used in making sulphuric acid/step 2502 + 02 2503
11 1 1 1 1 1 1 1 1 1 1 1 1 1 1 1 1 1 1 1
mattase breaks down mattose to simple sugar (glucose)
nickel used to hydrogenate unsaturated fats
Pt, Pd, Rh used in catalytic converters to breakdown NO2 in exhaus
<ol> <li>Describe what happens to the kinetic energy, potential energy and the total energy of reactant molecules as they approach each other.</li> </ol>
KE decreases PE increases TE remains constant.
I.d. Combain subura bassas materiales access for a constant to the
<ol> <li>Explain why a lower activation energy for a reaction leads to a greater reaction rate at a given temperature.</li> </ol>
molecular collisions of require as high an energy to be surressful
molecular collisions of require as high an energy to be successful so of a particular temp. a greater fraction of the molecules will have enough energy.  A small piece of this react with 2.0 M HCI to produce 12.0 mL of Hz was in 30.0 seconds.
15. A small piece of zigic react with 2.0 M HCl to produce 12.0 mL of H <sub>2</sub> gas in 30.0 seconds.
Calculate the rate of reaction
a) In mL of H <sub>2</sub> /second 12.0 mL H <sub>2</sub> = 0.400 mL H <sub>2</sub> /s 30.0 s
30.05 Answer 0.400 mL Hg/s
b) In moles of Hy/second
0.400 m L Hz × 1 L : 1 mol = 1.79 × 10 5 mol Hz/s Answer 1.79 × 10 5 mol Hz/s
1.79×10 5 mol Hz/s

# **Modeling Chemistry U7 Ws V2 0 Key**

**Minjie Lin** 

Modeling Chemistry U7 Ws V2 0 Key:

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