

Chemistry – Unit 5 Worksheet 1

Relative Abundance of the Elements

Relative Mass From Counts

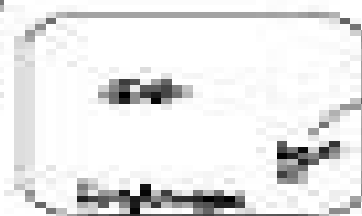
We can calculate the relative abundance of an element by dividing its mass by the mass of the most abundant isotope.

1. Equal numbers of atoms means the same number of nucleons as the most abundant isotope.
2. Number of nucleons = mass of the atom in atomic mass units.

Example: ${}^1_1\text{H}$, ${}^2_1\text{H}$, ${}^3_1\text{H}$, ${}^4_2\text{He}$, ${}^{12}_6\text{C}$, ${}^{16}_8\text{O}$

Example: ${}^1_1\text{H}$ is the most abundant isotope. Its mass is 1.007825 amu. The mass of the other isotopes is: ${}^2_1\text{H}$ is 2.014102 amu, ${}^3_1\text{H}$ is 3.016049 amu, ${}^4_2\text{He}$ is 4.002603 amu, ${}^{12}_6\text{C}$ is 12.000000 amu, ${}^{16}_8\text{O}$ is 15.994915 amu. The relative abundance of ${}^1_1\text{H}$ is 1.000000.

Hydrogen	1.007825
Helium	4.002603
Lithium	7.016003



Hydrogen	1.007825
Helium	4.002603
Lithium	7.016003

For example, the relative abundance of ${}^2_1\text{H}$ is 0.0156. This means that for every 100 atoms of ${}^1_1\text{H}$, there are 0.0156 atoms of ${}^2_1\text{H}$. The relative abundance of ${}^3_1\text{H}$ is 0.0001. This means that for every 1000 atoms of ${}^1_1\text{H}$, there are 0.0001 atoms of ${}^3_1\text{H}$.

Relative Mass From Composition

Many elements are made up of two or more isotopes. The relative abundance of the isotopes can be found by dividing the mass of the element by the mass of the most abundant isotope. The relative abundance of the isotopes can also be found by dividing the mass of the element by the mass of the most abundant isotope.

For example, the relative abundance of ${}^1_1\text{H}$ is 1.000000. The relative abundance of ${}^2_1\text{H}$ is 0.0156. The relative abundance of ${}^3_1\text{H}$ is 0.0001. The relative abundance of ${}^4_2\text{He}$ is 0.0003. The relative abundance of ${}^{12}_6\text{C}$ is 0.9893. The relative abundance of ${}^{16}_8\text{O}$ is 0.9999.



Modeling Chemistry U5 Ws1 V3 0

Marcel A. Müller



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