

AP Chemistry Unit 3 Review

Lewis Diagrams, IMFs, and Gases

Honors Lecture Notes: Unit 2, Unit 3 Lectures 4-7

Lewis Diagrams

Steps to creating a Lewis Diagram:

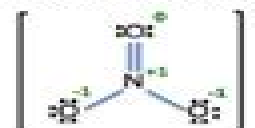
- 1) Count the number of valence electrons
- 2) Place the least electronegative element in the center and single bond all terminal atoms to it
- 3) Complete all octets around the terminal atoms
- 4) Add any extra valence electrons to the central atom*
- 5) If all valence electrons are used and the central atom is not full, create double or triple bonds

* Only atoms in period 3 or below can expand their octets

A resonance structure exists when there are multiple potential locations for a double bond. You draw structures with the double bond in each possible location, but in reality, the bonds are all identical and somewhere between single and double bonds in length and strength.

Formal Charge

To calculate the formal charge on an atom in a Lewis diagram, subtract the number of assigned electrons from the number of valence electrons for that atom. When counting assigned electrons, lone pairs count as two and bonds count as one.



Formal charges of greater than +1 or -1 do not exist in compounds. The less formal charge, the more stable (and thus, more likely) the compound is to exist. The total formal charge on all atoms in a polyatomic ion is equal to the overall charge on the ion.

VSEPR

Valence Shell Electron Pair Repulsion theory governs the shapes of molecules. A charge cloud is either a lone pair or any type of bond. Single/double/triple bonds all count as exactly one charge cloud. An outline of the shapes is below:

	0 lone pairs	1 lone pair	2 lone pairs	3 lone pairs
2 charge clouds	Linear			
3 charge clouds	Trigonal Planar	Bent		
4 charge clouds	Tetrahedral	Trigonal Pyramidal	Bent	
5 charge clouds	Trigonal Bipyramidal	See-Saw	T-Shaped	Linear
6 charge clouds	Octahedral	Square Pyramidal	Square Planar	

Shapes in the grey boxes are three-dimensional in nature.

Polarity

Molecules are polar if they are asymmetrical. This is usually the case when the central atom has lone pairs on it (exception: square planar and linear molecules are symmetrical). The most electronegative element (usually the terminal atom, except when H is terminal) gets a negative dipole, and the least electronegative element gets a positive dipole. If the central atom has no lone pairs, the molecule can still be polar if the terminal atoms are different.

Hybridization

Single bonds are made up of sigma bonds. Double bonds are one sigma and one pi bond, and triple bonds are one sigma and two pi bonds.

Central atoms with 2 charge clouds exhibit sp hybridization.
Central atoms with 3 charge clouds exhibit sp^2 hybridization.
Central atoms with 4 charge clouds exhibit sp^3 hybridization.

Modeling Chemistry Unit 3 Review Answers

Christian G. Meyer



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Nuclear Science Abstracts , 1970-05 *Energy Research Abstracts* , 1990 **Review**

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